

INORGANIC CHEMISTRY



CHEMICAL BONDING

IONIC BONDS

- Atoms with low ionization energy form cations (1-), and atoms with high electron affinity form anions (1-). Attraction between opposite charges of cations and anions creates an ionic bond.

THE LEWIS OCTET RULE: COVALENT BONDS

- Atoms other than hydrogen (stable with 2 valence electrons) will form bonds until they are surrounded by 8 valence electrons, and have achieved stable noble gas configuration.

The Lewis Dot Diagrams consist of one dot for each valence electron surrounding the symbol of the element.



BOND DISSOCIATION ENERGIES: CALCULATING ΔH

- $\Delta H = \sum \text{Bond energies of reactants} - \sum \text{Bond energies of products}$
- Example: $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- $\Delta H = 4(413.4) + 2(495.8) - 2(463) - 4(463.0)$
- $\Delta H = -890.4 \text{ kJ/mol}$ (exothermic reaction)
- Reaction is exothermic = energy (heat) flows out of system
- Endothermic = energy (heat) flows into system

Bond	Energy (kJ/mol)
H-H	436.4
H-O	463
H-Cl	431.9
C-H	413.4
C-C	347

PERCENT IONIC CHARACTER OF A CHEMICAL BOND

- Bonds with 50% or greater ionic character are ionic.
- Bonds with less than 50% ionic character are covalent.
- Transition from ionic to covalent bonding results in a reduction in electrical conductivity, melting point, and boiling point.



INTRODUCTION

- The study of compounds (group of carbon, except for CO , CO_2 , CS_2 , and compounds containing CN^- , CO_3^{2-} , and HCO_3^-)

INORGANIC COMPOUNDS

Name	Formula	Name	Formula
Alumina	Al_2O_3	Glauber's Salt	Na_2SO_4
Ammonia	NH_3	Cyanate	$\text{CaCN}_2 \cdot 2\text{H}_2\text{O}$
Anhydrous	SO_3	Heavy Water	D_2O
Baking soda	NaHCO_3	Litharge	PbO
Bauxite	$\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O} \cdot \text{H}_2\text{O}$	Salt or Magnesia	MgO
Calcite	CaCO_3	Carbide	CaC_2
Carbon	H_2CO_3	Silica	SiO_2
Cyanide	Na_3CN	Sulphur Lime	CaSO_4
Dry Ice	CO_2 (solid)	Table Salt	NaCl
Gypsum	$\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$	Washing Soda	$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$

COMMON INORGANIC IONS

Ammonium (NH_4^+)	Mercurous (Hg_2^{2+})	Nitride (N^{3-})
Ammonium (NH_4^+)	Mercuric (Hg^{2+})	Hydrogen Carbonate or Bicarbonate (HCO_3^-)
Barium (Ba^{2+})	Perchlorate (ClO_4^-)	Hydrogen Phosphate (HPO_4^{2-})
Calcium (Ca^{2+})	Silver (Ag^+)	Hydrogen Sulfate, Bisulfate (HSO_4^-)
Cadmium (Cd^{2+})	Sodium (Na^+)	Nitrate (NO_3^-)
Chromium (Cr^{3+})	Strontium (Sr^{2+})	Fluoride (F^-)
Chromium (Cr^{6+})	Tin (Sn), Stannous (Sn^{2+})	Selenate (SeO_4^{2-})
Cobalt (Co), Cobaltous (Co^{2+})	Zinc (Zn^{2+})	Selenide (Se^{2-})
Copper (Cu), Cuprous (Cu^+)	Bromide (Br^-)	Nitrite (NO_2^-)
Copper (Cu), Cupric (Cu^{2+})	Carbonate (CO_3^{2-})	Chloride (Cl^-)
Hydrogen (H^+)	Chlorate (ClO_3^-)	Permanganate (MnO_4^-)
Iron (Fe), Ferrous (Fe^{2+})	Chloride (Cl^-)	Formate (HCO_2^-)
Iron (Fe), Ferric (Fe^{3+})	Chromate (CrO_4^{2-})	Phosphate (PO_4^{3-})
Lead (Pb), Plumbic (Pb^{2+})	Cyanide (CN^-)	Sulfate (SO_4^{2-})
Lithium (Li^+)	Dichromate ($\text{Cr}_2\text{O}_7^{2-}$)	Sulfide (S^{2-})
Magnesium (Mg), Magnesian (Mg^{2+})	Ethylhydrogen Phosphate ($\text{C}_2\text{H}_5\text{HPO}_4^-$)	Sulfite (SO_3^{2-})
Magnesium (Mg^{2+})	Bromate (BrO_3^-)	Thiosulfate ($\text{S}_2\text{O}_3^{2-}$)

PAULI'S RULE

- Indicates that one orbital has the same energy, electron occupancy rules, and do not with parallel spins.
- Using 2p orbitals to demonstrate.

ORBITAL TYPES & SHAPES



PAULI'S EXCLUSION PRINCIPLE

- No 2 electrons in an atom can have the same 4 quantum numbers. No more than 2 electrons can occupy any 1 orbital.

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