

Introduction:

Water is an integral part of many ionic solids and such ionic solids are called *hydrates*. The water in these solids is called *waters of hydration*. A familiar example of a hydrate is plaster of Paris, which is the hemihydrate of calcium sulfate, $\text{CaSO}_4 \bullet \frac{1}{2} \text{H}_2\text{O}$. When water is added to plaster of Paris and the material is allowed to set, it is gradually transformed into a hard crystalline compound, calcium sulfate dihydrate, $\text{CaSO}_4 \bullet 2\text{H}_2\text{O}$. This is the material of plaster casts. The difference in composition between plaster of Paris and the plaster in casts is directly associated with the different degree of hydration of the calcium sulfate in the two cases.

The water of hydration is not as tightly bound in the hydrated crystal as the ions are. The water can usually be driven off by heating the crystals in a burner flame. The material that remains after the water has been removed is called the anhydrous salt.



An example of such a change is seen in the heating of calcium sulfate dihydrate.



In this experiment, you will heat several hydrates and observe the appearance of the reactants and products involved. A non-hydrated compound will also be heated to serve as a basis for comparison.

Objectives

1. Observe the effect of heat on hydrated ionic salts.
2. Verify the hydrate formula using stoichiometric methods

Procedure.

1. Obtain 2 test tubes and record the mass of each. Record
2. Place 2-3 grams of each substance into the appropriate test tube, mass the test tube and hydrate. Record
3. Using a test tube holder, holding the test tube nearly horizontal, gently heat the test tube. Note any changes in the appearance in the crystals.
4. Allow the test tube to cool and mass. Record
5. Repeat steps 3 & 4 until the mass of the ionic salts remains the same.
6. Repeat steps 2 -5 for the same ionic salt.

Data

Test Tube	Mass of Test Tube	Mass of Test Tube & Hydrate	Mass of Hydrate	Mass of Anhydrous	Mass of Water
1					
2					

Observations

Tube	Ionic Salt/hydrate	Observations
1		
2		

Hydrates Lab Answers

Karin Nielsen-Saines

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