

Chapter 1: Stoichiometry

Q4.

Sol: $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$

Since 1 mole of CaCO_3 produces 1 mole of CaO , So

1 mole of CaO will produce 1 mole of CaCO_3 .

$\therefore$ 1 mole of CaO $\rightarrow$ 1 mole of CaCO_3 .

Sol: $\text{BaCl}_2 \rightarrow \text{BaSO}_4$

Given:

5 moles of BaCl_2 contains 1 mole of Ba & 10 moles of Cl & each form 1 mole of BaSO_4 , $\therefore$ since 1 mole of BaSO_4 contains 1 mole of Ba

Sol: $2 \text{Na} \rightarrow \text{Na}_2\text{O}$

5 moles O_2 given

$\therefore$ 2 moles produces 1 mole of Na_2O

$\therefore$ 5 $\rightarrow$ $\frac{5}{2}$

$\therefore$ 5 $\rightarrow$ $\frac{5}{2} \times 2 = 5$ moles of Na_2O .

Q1.

Sol: $\text{Ca} + \frac{1}{2} \text{O}_2 \rightarrow \text{CaO}$ (balanced chemical equation)

Wt. of $\text{Ca} = 40 \text{ g/mole}$

So, moles of $\text{Ca} = \frac{80}{40} = 2 \text{ moles}$

From above balanced chemical equation: 1 moles of Ca reacts with $\frac{1}{2}$ moles of O_2 .

Wt. of O_2 with which 80g of Ca react $= \frac{1}{2} \times 32 = 16 \text{ g}$

Ans.

Q2.

Sol: $\text{NaCl} \rightarrow \text{NaCl}$

50 gm

If we have not balanced the chemical equation, then we have to apply P.M.C. rule or g-equivalent rule.

By P.M.C. rule (I solved)

Mass of Cl absorbed before reaction = Mass of Cl absorbed after reaction

Mass of moles of $\text{NaCl} = \text{Wt.} = 2 \times \text{mass of moles of } \text{Cl}_2 = 71 \text{ g}$

$\frac{50}{58.5} = 2 \times \text{moles of } \text{Cl}_2$

moles of $\text{Cl}_2 = \frac{50}{117}$

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