# 10 • Orbital Hybridization & Molecular Orbitals

#### PRACTICE TEST

- Which hybridization is associated with a steric number of 37
  - 2 Ce 98 18

- F-B-FI
- The molecule BrF<sub>3</sub> has a steric number of on the central atom?
  - **63** 3
- b) 4

- What is the hybridization of Br in BrF 1?

(d) sp'd SN = 5

b) sp

- e) sp
- How many equivalent sp<sup>3</sup>d orbitals are there?
  - m) 3

- do 6
- What type of hybridization is associated with a square planar molecular shape?
  - in b see !

do spid

by sp2

- ch sp
- What shape for electron pairs is associated with sp<sup>3</sup>d<sup>2</sup> hybridization?
  - ia) limear

- tetrahedrali
- b) square planar
- octahedral

- c) bent
- What hybridization is predicted for phosphorus in the PCI<sub>1</sub> molecule?
  - p)  $sp^2$

- Letreholal superindal

- A double bond contains sigma bond(s) and ( pi bond(s).
  - a) 0.2

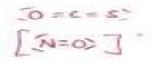
- c) 2,0 (d) 1,1
- What angle exists between orbitals in sp3d2 hybrid perhitals/2
  - 90.00 180.0\*

- 109.5
- 10. Which of the following elements is most likely to display sp<sup>3</sup>d hybridization? Su = 3
  - at oxygen
- bb- mitrogen

- c) phosphorus
- How many sigma (σ) and pi (π) electrons pairs are
  - a) four σ and zero π d) two σ and four π
  - b) three  $\sigma$  and two  $\pi$  e) one  $\sigma$  and three  $\pi$
  - (c) ) two o and two x
- 12. What is the hybridization of the oxygen atoms in
- (0-64 CH<sub>3</sub>OH and CO<sub>2</sub>, respectively? (0=C = 0)
- d-€-µ a) sp', sp'
- d) sp<sup>2</sup>, sp<sup>2</sup> ( Su = 3

- - All of the following species contain two π-bonds EXCEPT
    - a) SCN 16 a-
- b) CO IDe-
- c) H<sub>2</sub>CCO Hear





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