

CP Chemistry Unit 1 Test Review KEY

Problems and questions:

- Round each of the following numbers to 3 significant figures:
a) .09882 **.0988** b) 123,459 **123,000** c) 102.45 **102** d) 2.301×10^{-2} **2.30×10^{-2}**
- Calculate, use sig figs: $(12.22 \times 14.02 \times 120.1 \times .555) / 10.00 =$ **1140**
- A sample of metal having a mass of 12.122 g is added to a graduated cylinder containing 35.2 mL of water. The water level rises to 38.4 mL. Calculate the density, use correct significant figures.
$$d = \frac{12.122}{3.2} = 3.8 \text{ g/mL}$$
- Calculate the % error for an experiment where the student measured 3.45 g and the actual value was 3.00 g?
$$\% \text{ error} = \frac{(3.45 \text{ g} - 3.00 \text{ g})}{(3.00 \text{ g})} \times 100 = 15.0\%$$
- Write the symbol for the following ions: Mg losing 2 e⁻, Al losing 3 e⁻, N gaining 3 e⁻, S gaining 2 e⁻
Mg²⁺, Al³⁺, N³⁻, S²⁻
- a) 2 b) 3 c) 12 d) 5
- List three points of Dalton's atomic theory.
A. Elements made of tiny particles, called atoms.
B. Atoms of a given element are identical; the atoms of different elements are different.
C. In chemical reactions, atoms are combined, separated, or rearranged. Atoms cannot be subdivided, created, or destroyed.
D. Atoms combine with atoms, in small whole numbers, to form compounds.
- Describe Thompson's experiment with the cathode ray tube, what did he discover?
Cathode ray tube consists of a cathode (negative) and anode (positive) side. The "ray" could be bent with magnetic or electric fields. The "ray" bent toward a positive EM field and away from a negative EM field. Discovered electron and charge to mass ratio of electron.
- Describe Millikan's Oil Drop experiment. What exactly did he determine?
He suspended a micro oil droplet in an electric field and was able to determine the actual charge of an electron to be $-1.6 \times 10^{-19} \text{ C}$. Also, determined the mass of an electron.
- Explain Rutherford's Gold Foil experiment. What were three observations and three conclusions?
He fired alpha particles at gold foil and observed: 1) most went straight through, 2) small number showed a slight deflection, 3) very few bounced straight back. He concluded: 1) mass of atom is in the nucleus, 2) nucleus is positively charged, 3) atom mostly empty space.
- What is an isotope? **Atoms of the same element that differ by # of electrons.**
- Write the symbol of the most abundant copper isotope if the average atomic mass is 63.55? **$^{64}_{29}\text{Cu}$**
- Write the nuclear symbol for an isotope with Z=40, and 45 neutrons? **$^{85}_{40}\text{Zr}$**
- Consider the diagram and match the letters with the appropriate particle
A= electron, B= neutron, C= proton
- Place the forms of hydrogen (H, H⁺, H⁻, H₂) in the table.



Atom	Cation	Anion	Molecule
H	H⁺	H⁻	H₂

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R Sanford



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Table of Contents Modeling Chemistry Unit 1 Review Answers

1. Understanding the eBook Modeling Chemistry Unit 1 Review Answers
 - The Rise of Digital Reading Modeling Chemistry Unit 1 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 1 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 1 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 1 Review Answers
 - Personalized Recommendations

- Modeling Chemistry Unit 1 Review Answers User Reviews and Ratings
- Modeling Chemistry Unit 1 Review Answers and Bestseller Lists
- 5. Accessing Modeling Chemistry Unit 1 Review Answers Free and Paid eBooks
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 - Modeling Chemistry Unit 1 Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 1 Review Answers
 - Highlighting and Note-Taking Modeling Chemistry Unit 1 Review Answers
 - Interactive Elements Modeling Chemistry Unit 1 Review Answers
- 8. Staying Engaged with Modeling Chemistry Unit 1 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 1 Review Answers
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 1 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 1 Review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 1 Review Answers
 - Setting Reading Goals Modeling Chemistry Unit 1 Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 1 Review Answers
 - Fact-Checking eBook Content of Modeling Chemistry Unit 1 Review Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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