At a high temperature, 0.50 mol of HBr was placed in a 1.0 L container and allowed to decompose according to the reaction:

$$2HBr(g) \rightleftharpoons H_{2(g)} + Br_{2(g)}$$

At equilibrium the [Br₂] was measured to be 0.13 M. What is K_{eq} for this reaction at this temperature?

| | 2HBr = | = H2 + | Br2 |
|-----|--------|--------|-------|
| [x] | 0.50 | 0 | 0 |
| [C] | -0-26 | +0.13 | +0.13 |
| IE] | 0.24 | 0.13 | 0.13 |

| Keg = | IH2][Br]2 | = (0.13)2 | =0.29 |
|-------|---|-----------------|-------|
| | 100001100000000000000000000000000000000 | 150000000000000 | |

Answer Keg = 0.29

 When 1.0 mol of NH_{3(g)} and 0.40 mol of N_{2(g)} are placed in a 5.0 L vessel and allowed to reach equilibrium at a certain temperature, it is found that 0.78 mol of NH₃ is present. The reaction is:

2NH3(g) = 3H2(g) + N2(g)

Initial [NH3] = 1.0 mol = 0.20 M Initial [N2] = 0.40 mol = 0.080 M

Equil [NH3] = 0.78 mol = 0.156 M 2NH2 = 3 H2 + N2

| | aNH3 & | = 3 Na | + N2 |
|------|--------|--------|--------|
| [I] | 0.20 | 0 | 0.080 |
| EC3 | -0.044 | +0.066 | +0.022 |
| LEJ. | 0.156 | 0.066 | 0.102 |

a) Calculate the equilibrium concentrations of all three species.

[NH₃] - 0.16 M [H₂] - 0.066 M [N₂] - 0.10 M

b), Calculate the value of the equilibrium constant at this temperature. (use unrounded cones, then round to 2503)

 $\text{Keq} = \frac{\text{EH}_2 \text{J}^3 \text{EN}_2 \text{J}}{\text{ENH}_3 \text{J}^2} = \frac{(0.066)^3 (0.102)}{(0.156)^2} = 0.0012$ Answ

Answer 1.2×10-3

c) How many moles of H₂ are present at equilibrium?

0.066 M x 5.0 L = 0.33 mal

Answer 0.33 mol

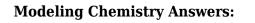
d) How many moles of N₂ are present at equilibrium?

0.102 M x 5.0 L = 0.51 mal

Answer 0.51 mol

Modeling Chemistry Answers

Douw G. Steyn, S. T. Rao



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