

Acids & Bases Review

1. Name the following compounds as acids and circle which are weak acids:



nitric acid



sulfuric acid



hydrofluoric acid



carbonic acid



acetic acid



phosphoric acid

2. Write the formulas for the following bases:

a) calcium hydroxide



b) aluminum hydroxide



c) ammonium hydroxide



d) lithium hydroxide



3. Complete the following table:

	Acids	Bases
Definition	A substance that increases the concentration of H^+ (or H_3O^+) when added to water. or A substance that acts as a proton (H^+) donor.	A substance that increases the concentration of OH^- when added to water. or A substance that acts as a proton (H^+) acceptor.
Electrolyte in aqueous solution?	H^+	OH^-
pH	less than 7	greater than 7
Taste	sour	bitter
Litmus turns	blue turns red	red turns blue
Which reacts with metals to form H_2 gas?	yes	no

Conjugate acid-base pairs:

4. Identify the acid, the base, the conjugate acid and the conjugate base in the reactions:



5. An acid will form the conjugate base and a base will form the conjugate acid.

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