

HYDROLYSIS OF SALTS AND pH OF BUFFER SOLUTIONS

INTRODUCTION

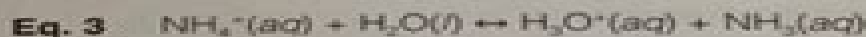
The purpose of this experiment is to determine the pH of various concentrations of acids, bases, salts, and buffer solutions. The data will be used to determine the acid-base nature in the hydrolysis of several salts and behavior of buffer solutions. Recall that salts are the product of neutralization reactions between acids and bases, generically shown below with an acid, HA, and base, B.



As nearly all salts are strong electrolytes, they exist in solution as ions. Many of these ions then can react with water to produce an acidic or basic solution. This reaction of salts with water is called a hydrolysis reaction. An anion, such as the generic A^- , is the conjugate base of the weak acid HA. These anions (A^-) will then react with water to generate OH^- .



Cations such as NH_4^+ or highly charged metals like Fe^{3+} come from weak bases and react with water to form H^+ (H_3O^+ , hydronium) ions.



HYDROLYSIS OF ANIONS: BASIC SOLUTIONS

First, let us focus on the hydrolysis of anions. For example, the nitrite ion (NO_2^-), which will react with water (like a weak base, for example NH_3) to generate hydroxide and thus a basic solution (pH greater than 7):



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Mike Jess



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Chemical Abstracts ,1910 **Laboratory Experiments for General Chemistry** Harold R. Hunt,Toby F. Block,1994
Merrill Chemistry-Lab.Manual Smoot,1994-07 U.S. Government Research Reports ,1964 **Scientific and Technical Aerospace Reports** ,1995 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database **General Chemistry** Hunt,1976-06-07 **Nuclear Science Abstracts** ,1974-06 **Bulletin** ,1977
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