

Hewlett-Packard Financial Performance Assessment

Exercise 10.17 has three similar following problems. Use all appropriate and proper scientific units. Use scientific notation for answers (e.g., 1.23×10^4).

1. Find the standard enthalpy change for the reaction $\text{C}_{10}\text{H}_{18}\text{O} + 15\text{O}_2 \rightarrow 10\text{CO}_2 + 9\text{H}_2\text{O}$.
 Given that $\Delta H_f^\circ(\text{C}_{10}\text{H}_{18}\text{O}) = 155 \text{ kJ/mol}$ and $\Delta H_f^\circ(\text{C}_{10}\text{H}_{18}) = 150 \text{ kJ/mol}$.
 and $\Delta H_f^\circ(\text{H}_2\text{O}) = -285 \text{ kJ/mol}$ and $\Delta H_f^\circ(\text{CO}_2) = -393 \text{ kJ/mol}$.

a. $\Delta H_{\text{rxn}}^\circ = 155 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 150 \text{ kJ/mol}$ $\Delta H_f^\circ = -285 \text{ kJ/mol}$
 $\Delta H_f^\circ(\text{C}_{10}\text{H}_{18}\text{O}) = 155 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{C}_{10}\text{H}_{18}) = 150 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{H}_2\text{O}) = -285 \text{ kJ/mol}$

b. $\Delta H_{\text{rxn}}^\circ = 155 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 150 \text{ kJ/mol}$ $\Delta H_f^\circ = -285 \text{ kJ/mol}$

2. The standard enthalpy change for the formation of aluminum oxide and aluminum nitride are $\Delta H_f^\circ(\text{Al}_2\text{O}_3) = -1675 \text{ kJ/mol}$ and $\Delta H_f^\circ(\text{Al}_3\text{N}) = -170 \text{ kJ/mol}$.
 $\Delta H_f^\circ(\text{Al}_2\text{O}_3) = -1675 \text{ kJ/mol}$ and $\Delta H_f^\circ(\text{Al}_3\text{N}) = -170 \text{ kJ/mol}$.
 Calculate the standard enthalpy change for the reaction $2\text{Al}_3\text{N} + 3\text{Al}_2\text{O}_3 \rightarrow 6\text{Al} + 3\text{N}_2$.

a. $\Delta H_{\text{rxn}}^\circ = 1675 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 170 \text{ kJ/mol}$ $\Delta H_f^\circ = -1675 \text{ kJ/mol}$
 $\Delta H_f^\circ(\text{Al}_2\text{O}_3) = -1675 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{Al}_3\text{N}) = -170 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{Al}) = 0 \text{ kJ/mol}$

b. $\Delta H_{\text{rxn}}^\circ = 1675 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 170 \text{ kJ/mol}$ $\Delta H_f^\circ = -1675 \text{ kJ/mol}$
 $\Delta H_f^\circ(\text{Al}_2\text{O}_3) = -1675 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{Al}_3\text{N}) = -170 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{Al}) = 0 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}}^\circ = 1675 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 170 \text{ kJ/mol}$ $\Delta H_f^\circ = -1675 \text{ kJ/mol}$

3. The gas-phase standard enthalpy of formation of carbon monoxide, carbon dioxide, and hydrogen, listed from top to bottom are $\Delta H_f^\circ(\text{CO}) = 110 \text{ kJ/mol}$, $\Delta H_f^\circ(\text{CO}_2) = -393 \text{ kJ/mol}$, and $\Delta H_f^\circ(\text{H}_2) = 0 \text{ kJ/mol}$.

Calculate the standard enthalpy change for the reaction from the following chemical equation:

$$2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$$

a. $\Delta H_{\text{rxn}}^\circ = 110 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 110 \text{ kJ/mol}$ $\Delta H_f^\circ = 110 \text{ kJ/mol}$
 $\Delta H_f^\circ(\text{CO}) = 110 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{O}_2) = 0 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{CO}_2) = -393 \text{ kJ/mol}$

b. $\Delta H_{\text{rxn}}^\circ = 110 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 110 \text{ kJ/mol}$ $\Delta H_f^\circ = 110 \text{ kJ/mol}$
 $\Delta H_f^\circ(\text{CO}) = 110 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{O}_2) = 0 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{CO}_2) = -393 \text{ kJ/mol}$
 $\Delta H_f^\circ(\text{H}_2) = 0 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{H}_2\text{O}) = -285 \text{ kJ/mol}$ $\Delta H_f^\circ = 110 \text{ kJ/mol}$

c. $\Delta H_{\text{rxn}}^\circ = 110 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 110 \text{ kJ/mol}$ $\Delta H_f^\circ = 110 \text{ kJ/mol}$
 $\Delta H_f^\circ(\text{CO}) = 110 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{O}_2) = 0 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{CO}_2) = -393 \text{ kJ/mol}$
 $\Delta H_f^\circ(\text{H}_2) = 0 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{H}_2\text{O}) = -285 \text{ kJ/mol}$ $\Delta H_f^\circ = 110 \text{ kJ/mol}$

d. $\Delta H_{\text{rxn}}^\circ = 110 \text{ kJ/mol}$ $\Delta H_{\text{rxn}}^\circ = 110 \text{ kJ/mol}$ $\Delta H_f^\circ = 110 \text{ kJ/mol}$

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