

Heat Problems I

Standards:

- 3.1.10 B Describe the concepts of models as a way to predict and understand science and technology.
- 3.4.10 B Analyze energy sources and transfers of heat.

Specific Heat Capacities of some Common Substances

Substance	Cp (J/g°C)	Substance	Cp (J/g°C)
Water (l)	4.184	Iron (s)	0.45
Water (s)	2.03	Mercury (l)	0.14
Water (g)	2.0	Carbon (s)	0.71
Aluminum (s)	0.89	Silver (s)	0.24
		Gold (s)	0.13

$$1\text{Cal} = 4.184\text{Joules}$$

$$1\text{kJ} = 1000\text{J}$$

Problems:

- Convert the following number of calories into joules. (Use D.A. and sig figs.)
 - 100.0 cal = _____ J
 - 1.00×10^3 cal = _____ J
- Convert the following numbers of joules (J) into kilojoules (kJ). (Use D.A. and sig figs.)
 - 243,000 J = _____ kJ
 - 0.251 J = _____ kJ

Heat And Temperature Answer Sheet

Jianjun Gao



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