Recult's law for a volatile liquid - panered vapous pressure of 1 P - panelel vapour pressure of 2 p - vapour pressure of pane 2 p, - p, x, , p, - p, x, p" vapous partisure of puse 2 Total pressure p = p. + p. = p. x, + p. x, Ideal Solution 1) Rapults law is obeyed Relative lowening of vapoun pressure 2.) Amix H = 0 5) Amx V - 0 Elevation of boiling point | ATb - Kbm Ka - ebulliscopic constant Kb X W. X 1000 Depression of frezzing point Ky . Chyoscopic constant ATr - Krm ATE - KEX W. X 1000 W. . Mass of Solvent Osmotic pressure (TT) TT = W. " Mass of Solute M, . Molan mass of solvent M. . Molan mass of solute C - Molanity T - Tempenature TT - W. X RXT M. X V Van't Hoff factor (i) Degree of Disassociation - Normal molar mass 4-1 m - mumber of ions Abnormal molar mass Observed colligative property NCERT Calculated colligative property Total no of moles of particles after association / dissassociation no of moles of particles before association / dissassociation Modified form of Colligative properties 2) ATh - 1 Kh m 1) ATe = iKem 4) IT = in, RT/V

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