

Chemistry – Unit 5 Worksheet 2 Empirical and Molecular Formulas

Show all your work, show all steps, show all calculations. Do not forget to show units and label your answers.

1. Find the empirical formula of a compound containing 33.9 g of bromine and 4.7 g of nitrogen.

$$\text{mole Br} = \frac{33.9 \text{ g Br}}{79.90 \text{ g/mol}} = 0.424 \text{ mole Br}$$

1.00 mole

$$\text{mole N} = \frac{4.7 \text{ g N}}{14.01 \text{ g/mol}} = 0.335 \text{ mole N}$$



2. What is the empirical formula of a hydrocarbon compound if 1.00 g of this compound contains 0.75 g of carbon and 0.25 g of hydrogen?

$$\text{mole C} = \frac{0.75 \text{ g C}}{12.01 \text{ g/mol}} = 0.062 \text{ mole C}$$

$$\text{mole H} = \frac{0.25 \text{ g H}}{1.008 \text{ g/mol}} = 0.25 \text{ mole H}$$

1.00 mole

$$\text{CH}_4 = \frac{0.25}{0.062} = 4.0 \text{ mole H}$$



3. A compound was identified and determined to contain 40.0 g of nitrogen, 10.0 g of hydrogen, and 10.0 g of oxygen. What is the empirical formula of this compound?

$$\text{mole N} = \frac{40.0 \text{ g N}}{14.01 \text{ g/mol}} = 2.86 \text{ mole N}$$

$$\text{mole H} = \frac{10.0 \text{ g H}}{1.008 \text{ g/mol}} = 9.92 \text{ mole H}$$

1.00 mole

$$\text{mole O} = \frac{10.0 \text{ g O}}{16.00 \text{ g/mol}} = 0.625 \text{ mole O}$$



4. A compound composed of hydrogen and oxygen is found to contain 10.0 g of hydrogen and 10.0 g of oxygen. What is the value of the molar mass of this compound if 4.0 g of this compound contains 1.0 g of hydrogen?

$$\text{mole H} = \frac{1.0 \text{ g H}}{1.008 \text{ g/mol}} = 0.99 \text{ mole H}$$

$$\text{mole O} = \frac{1.0 \text{ g O}}{16.00 \text{ g/mol}} = 0.062 \text{ mole O}$$

1.00 mole hydrogen = 1.00 mole H₂O

1.00 mole oxygen = 0.50 mole O₂

0.50 mole H₂O = 0.50 mole H₂O

0.50 mole H₂O = 0.50 mole H₂O

0.50 mole H₂O = 0.50 mole H₂O

Modeling Chemistry Unit 5 Review Answers

M Carnoy



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Table of Contents Modeling Chemistry Unit 5 Review Answers

1. Understanding the eBook Modeling Chemistry Unit 5 Review Answers
 - The Rise of Digital Reading Modeling Chemistry Unit 5 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 5 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 5 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 5 Review Answers
 - Personalized Recommendations
 - Modeling Chemistry Unit 5 Review Answers User Reviews and Ratings
 - Modeling Chemistry Unit 5 Review Answers and Bestseller Lists

5. Accessing Modeling Chemistry Unit 5 Review Answers Free and Paid eBooks
 - Modeling Chemistry Unit 5 Review Answers Public Domain eBooks
 - Modeling Chemistry Unit 5 Review Answers eBook Subscription Services
 - Modeling Chemistry Unit 5 Review Answers Budget-Friendly Options
6. Navigating Modeling Chemistry Unit 5 Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 5 Review Answers Compatibility with Devices
 - Modeling Chemistry Unit 5 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 5 Review Answers
 - Highlighting and Note-Taking Modeling Chemistry Unit 5 Review Answers
 - Interactive Elements Modeling Chemistry Unit 5 Review Answers
8. Staying Engaged with Modeling Chemistry Unit 5 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 5 Review Answers
9. Balancing eBooks and Physical Books Modeling Chemistry Unit 5 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 5 Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Unit 5 Review Answers
 - Setting Reading Goals Modeling Chemistry Unit 5 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Unit 5 Review Answers
 - Fact-Checking eBook Content of Modeling Chemistry Unit 5 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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