

Unit 6 – Worksheet 3

Ionic Compounds

Properties

Basic structural unit

1. Give the name of the following simple binary ionic compounds.

- a. Na_2O
- b. K_2S
- c. MgCl_2
- d. CaBr_2
- e. BaI_2
- f. Al_2S_3
- g. CsBr
- h. AgF

2. Give the name of the following simple binary ionic compounds.

- a. Na_3N
- b. K_2O
- c. AgBr
- d. MgI_2
- e. SrO

3. Write the formula for the following binary ionic compounds.

- a. lithium bromide
- b. sodium iodide
- c. silver sulfide
- d. cesium oxide
- d. beryllium iodide
- f. barium hydride
- g. aluminum fluoride
- h. potassium oxide

4. Write the formula for these ionic substances.

- a. silver oxide
- b. aluminum sulfide
- c. sodium nitride
- d. barium chloride
- e. strontium hydride

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Table of Contents Modeling Chemistry U6 Ws 5 V1 0

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 - The Rise of Digital Reading Modeling Chemistry U6 Ws 5 V1 0
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U6 Ws 5 V1 0
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U6 Ws 5 V1 0
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U6 Ws 5 V1 0
 - Personalized Recommendations
 - Modeling Chemistry U6 Ws 5 V1 0 User Reviews and Ratings
 - Modeling Chemistry U6 Ws 5 V1 0 and Bestseller Lists
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- Modeling Chemistry U6 Ws 5 V1 0 eBook Subscription Services
- Modeling Chemistry U6 Ws 5 V1 0 Budget-Friendly Options
- 6. Navigating Modeling Chemistry U6 Ws 5 V1 0 eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U6 Ws 5 V1 0 Compatibility with Devices
 - Modeling Chemistry U6 Ws 5 V1 0 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U6 Ws 5 V1 0
 - Highlighting and Note-Taking Modeling Chemistry U6 Ws 5 V1 0
 - Interactive Elements Modeling Chemistry U6 Ws 5 V1 0
- 8. Staying Engaged with Modeling Chemistry U6 Ws 5 V1 0
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U6 Ws 5 V1 0
- 9. Balancing eBooks and Physical Books Modeling Chemistry U6 Ws 5 V1 0
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U6 Ws 5 V1 0
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U6 Ws 5 V1 0
 - Setting Reading Goals Modeling Chemistry U6 Ws 5 V1 0
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U6 Ws 5 V1 0
 - Fact-Checking eBook Content of Modeling Chemistry U6 Ws 5 V1 0
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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