

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<hr/> 58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<hr/> 84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<hr/> 159.6

Modeling Chemistry U5 Wsvkey

Yicheng Fang



Modeling Chemistry U5 Wsvkey:

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Table of Contents **Modeling Chemistry U5 Wsvkey**

1. Understanding the eBook **Modeling Chemistry U5 Wsvkey**
 - The Rise of Digital Reading **Modeling Chemistry U5 Wsvkey**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Modeling Chemistry U5 Wsvkey**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Modeling Chemistry U5 Wsvkey**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Modeling Chemistry U5 Wsvkey**
 - Personalized Recommendations
 - **Modeling Chemistry U5 Wsvkey** User Reviews and Ratings
 - **Modeling Chemistry U5 Wsvkey** and Bestseller Lists

5. Accessing Modeling Chemistry U5 Wsvkey Free and Paid eBooks
 - Modeling Chemistry U5 Wsvkey Public Domain eBooks
 - Modeling Chemistry U5 Wsvkey eBook Subscription Services
 - Modeling Chemistry U5 Wsvkey Budget-Friendly Options
6. Navigating Modeling Chemistry U5 Wsvkey eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U5 Wsvkey Compatibility with Devices
 - Modeling Chemistry U5 Wsvkey Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U5 Wsvkey
 - Highlighting and Note-Taking Modeling Chemistry U5 Wsvkey
 - Interactive Elements Modeling Chemistry U5 Wsvkey
8. Staying Engaged with Modeling Chemistry U5 Wsvkey
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U5 Wsvkey
9. Balancing eBooks and Physical Books Modeling Chemistry U5 Wsvkey
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U5 Wsvkey
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry U5 Wsvkey
 - Setting Reading Goals Modeling Chemistry U5 Wsvkey
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry U5 Wsvkey
 - Fact-Checking eBook Content of Modeling Chemistry U5 Wsvkey
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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