

Chemistry – Unit 5 Worksheet 1

Relative Abundance of the Elements

Relative Mass From Counts

We can estimate relative abundances by counting all particles in the substance of a known composition and mass.

1. Equal numbers of gases occupy the same volume at the same pressure and temperature. **Avogadro's Hypothesis**
2. Number of molecules in a gas is proportional to its mass (molar mass).

Ex. H_2 , O_2 , N_2 , Ar , CO_2 , He

Using the periodic table, we can determine approximately how the density of gases are related to composition and pressure. In 1911, physical chemists began measuring gas densities. Measurements in 1919 showed that the composition of air (which is mostly nitrogen and oxygen) is 78% nitrogen, 21% oxygen, and 1% argon.

Oxygen	Hydrogen
20 g O_2	2 g H_2
1 Liter	1 Liter



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20 g O_2	2 g H_2
1 Liter	1 Liter

Physicists eventually determined ways to determine the relative masses of all the elements using this technique. Measurements of the density of the gases at the same pressure and volume are difficult, it was impossible. However, we can apply the Avogadro's Hypothesis concept and use the periodic table to determine relative abundances.

Relative Mass From Composition

Many substances contain known ratios of different types of atoms. For example, water is 11% hydrogen and 89% oxygen. This is a constant ratio. The composition of a substance is constant. The composition of a substance is constant.

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Peipei Pang



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