

Chemistry Unit 8 Review

1. The gas ethane, C_2H_6 , reacts with oxygen gas to produce the CO_2 oxide and water vapor.
 If 2.150 grams of CO_2 were produced using this chemical reaction, what mass of the gas water would be produced?



$$\text{Answer: } \underline{1.140 \text{ g H}_2\text{O}} \quad \text{Conversion: } \underline{\frac{18 \text{ g H}_2\text{O}}{44 \text{ g CO}_2}}$$

2. Magnesium nitrate, $\text{Mg}(\text{NO}_3)_2$, decomposes when heated to produce magnesium nitride and oxygen gas.
 If 2.14 grams of $\text{Mg}(\text{NO}_3)_2$ were heated, what mass of oxygen gas would be produced?



$$\text{Answer: } \underline{0.900 \text{ g O}_2} \quad \text{Conversion: } \underline{\frac{2 \text{ Mg}(\text{NO}_3)_2}{9 \text{ O}_2}} = \underline{\frac{2 \text{ Mg}(\text{NO}_3)_2}{9 \text{ O}_2}} = \underline{\frac{2(148 \text{ g Mg}(\text{NO}_3)_2)}{9(32 \text{ g O}_2)}} = \underline{\frac{2(148)}{9(32)}} = \underline{1.25 \text{ g O}_2/\text{g Mg}(\text{NO}_3)_2}$$

3. What mass of CO_2 was used up in the reaction with an excess of oxygen gas if 24.2 g of carbon monoxide is formed? $2 \text{ CO} + \text{O}_2 \rightarrow 2 \text{ CO}_2$



$$\text{Answer: } \underline{24.2 \text{ g CO}_2} \quad \text{Conversion: } \underline{\frac{2 \text{ mol CO}}{2 \text{ mol CO}_2}} = \underline{\frac{2(28 \text{ g CO})}{2(44 \text{ g CO}_2)}} = \underline{\frac{56 \text{ g CO}}{44 \text{ g CO}_2}} = \underline{1.27 \text{ g CO}_2/\text{g CO}}$$

4. The reaction between 12.8 g of CH_4 and excess oxygen gas produces 34.8 g of CO_2 gas and some water. Determine the percent yield.



$$\text{Answer: } \underline{50\%} \quad \text{Conversion: } \underline{\frac{12 \text{ g CH}_4}{44 \text{ g CO}_2}} = \underline{\frac{12 \text{ g CH}_4}{44 \text{ g CO}_2}} = \underline{\frac{12(16 \text{ g CH}_4)}{44(44 \text{ g CO}_2)}} = \underline{\frac{192}{1936}} = \underline{0.10 \text{ g CO}_2/\text{g CH}_4}$$

Modeling Chemistry Unit 8 Key

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