

## Chemistry Unit 8 Review

1. The salt chlorate,  $\text{NaClO}_3$ , reacts with oxygen gas to produce a few (2) solid acid chloride oxides.  
 If 5.152 grams of  $\text{Cl}_2$  were consumed using this chemical reaction, what mass of the salt would need to be produced?



$$5.152 \text{ g Cl}_2 \times \frac{1 \text{ mol}}{70.91 \text{ g}} = 0.07265 \text{ mol Cl}_2 \times \frac{2 \text{ mol NaClO}_3}{1 \text{ mol Cl}_2} = 0.1453 \text{ mol}$$

2. Magnesium nitrate,  $\text{Mg}(\text{NO}_3)_2$ , decomposes when heated to produce magnesium nitride and oxygen gas.  
 If 4.32 grams of  $\text{Mg}(\text{NO}_3)_2$  were heated to form this, what mass of oxygen gas would be produced?



$$4.32 \text{ g Mg}(\text{NO}_3)_2 \times \frac{1 \text{ mol}}{148.31 \text{ g}} = 0.02913 \text{ mol Mg}(\text{NO}_3)_2 \times \frac{12 \text{ mol O}_2}{2 \text{ mol Mg}(\text{NO}_3)_2} = 0.1748 \text{ mol O}_2 \times \frac{32.00 \text{ g}}{1 \text{ mol}} = 5.59 \text{ g O}_2$$

3. What mass of  $\text{Cl}_2$  was used up in the reaction with an excess of oxygen gas if 24.2 g of sodium chloride is formed?  $2 \text{NaCl} \rightarrow 2 \text{Na} + \text{Cl}_2$



$$24.2 \text{ g NaCl} \times \frac{1 \text{ mol}}{58.44 \text{ g}} = 0.4141 \text{ mol NaCl} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol NaCl}} = 0.2071 \text{ mol Cl}_2 \times \frac{70.91 \text{ g}}{1 \text{ mol}} = 14.7 \text{ g Cl}_2$$

4. The reaction between 11.8 g of  $\text{CH}_4$  and excess oxygen gas produces 11.8 g of  $\text{CO}_2$  gas and some water. Determine the percent yield.



$$11.8 \text{ g CH}_4 \times \frac{1 \text{ mol}}{16.04 \text{ g}} = 0.7356 \text{ mol CH}_4 \times \frac{1 \text{ mol CO}_2}{1 \text{ mol CH}_4} = 0.7356 \text{ mol CO}_2 \times \frac{44.01 \text{ g}}{1 \text{ mol}} = 32.58 \text{ g CO}_2$$

# Modeling Chemistry Unit 8 Key

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