

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out what products may be made. What are the possibilities and how do you do it? The next 4 problems cover typical questions.

1. Chlorine reacts with bromine in a reaction of gaseous molecules. $\text{Br}_2 + \text{Cl}_2$ say that, mixed chlorobromine gas produces gaseous chlorine. $\text{Br}_2 + \text{Cl}_2$ say, what is the balanced equation?

Equation:	$\text{Br}_2 + \text{Cl}_2 \rightarrow 2\text{BrCl}$
Reactants:	2 Br 2 Cl 0
Products:	0 Br 0 Cl 2 Br Cl
Balance:	0 0 0 0

Equation
balanced

2. Silver reacts with hydrogen chloride. Silver reacts with hydrogen chloride solution, producing silver chloride and hydrogen gas. This is exactly what we get when silver reacts with hydrogen chloride solution. How much silver chloride is produced if 100 g of silver reacts?

Equation:	$\text{Ag} + \text{HCl} \rightarrow \text{AgCl} + \text{H}_2$
Reactants:	100 0 0
Products:	0 0 100 0
Balance:	0 0 0 0

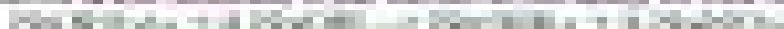
Equation
not balanced

3. Silver, Ag , reacts with hydrogen gas to produce silver chloride. How much silver chloride is produced if 100 g of silver reacts with hydrogen gas? (Silver is 107.87 g/mol, chlorine is 35.45 g/mol.)

Equation:	$\text{Ag} + \text{H}_2 \rightarrow \text{AgCl}$
Reactants:	100 0 0
Products:	0 0 100 0
Balance:	0 0 0 0

Equation is not balanced. $\text{Ag} + \text{H}_2 \rightarrow \text{AgCl}$ is not balanced.

4. Chlorine reacts with silver to produce silver chloride. How much silver chloride is produced if 100 g of silver reacts with chlorine? (Silver is 107.87 g/mol, chlorine is 35.45 g/mol.)



Equation is balanced.

Reactants:	100 0
Products:	0 0
Balance:	0 0

Equation is balanced. $2\text{Ag} + \text{Cl}_2 \rightarrow 2\text{AgCl}$ is balanced.

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A Gutmann



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