

Lab: Formula of a Chloride

Name _____

Introduction:

In this experiment we will attempt to determine the formula of the compound zinc chloride by measuring masses involved in the reaction of zinc with hydrochloric acid.

Safety Rules:

- Wear your safety glasses at all times during this lab.
- Do not drink from/been and make off joculars with strong odors.
- Wash your hands thoroughly before leaving the lab area.

Procedure:

- 1) Determine the mass of a clean, dry, 100 mL beaker to the nearest 0.001 g and record this mass in the data table.
- 2) Obtain a small amount of powdered zinc metal, and place it in the beaker. Then determine the mass of the beaker plus zinc and record in the data table.
- 3) Using a graduated cylinder to measure, add 10 mL of 10M HCl. Using a hot plate, gently heat the mixture until all of the zinc has dissolved.
- 4) Using a Bunsen burner, heat the solution until all of the water and excess HCl have been boiled away. Stop the heating as the last bit of liquid disappears. Do not continue heating after it has boiled to dryness as the compound will begin to decompose.
- 5) Cool for several minutes and then determine the mass of the beaker and contents.

Data:

mass of beaker	
mass of beaker + zinc	
mass of beaker + zinc chloride compound	

Calculations: Clearly label all calculations, and pay attention to sig figs!

1) Determine the mass of zinc used.

2) Determine the mass of chlorine that combined with the zinc.

3) Determine the moles of zinc used.

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D Keegan



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