

### Question 30

A student performs a titration with a base as the titrant and an acid as the analyte. He finds that the pH at the equivalence point is 8.3, and concludes that he was adding NaOH to HF. Which of the following observations made during the titration does NOT support his conclusion?

☒ A. The  $[H^+]$  concentration in the experimental vessel was higher before the titration began than at the equivalence point

You did not select. This answer is incorrect

☒ B. The solution, at the equivalence point, was found to be a conductor of electricity

You did not select. This answer is incorrect

☒ C. To reach the equivalence point, 10 mL of base (0.5 M) was added to 5 mL of acid (0.25 M)

You did not select. This answer is correct

☒ D. When the student titrates past the equivalence point, the solution reaches a pOH of 1

☒ This answer is incorrect

SEE EXPLANATION

# Kaplan Mcat Full Length 1 Answers

**JL Elias**



## **Kaplan Mcat Full Length 1 Answers:**

The book delves into Kaplan Mcat Full Length 1 Answers. Kaplan Mcat Full Length 1 Answers is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Kaplan Mcat Full Length 1 Answers, encompassing both the fundamentals and more intricate discussions.

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  - Chapter 1: Introduction to Kaplan Mcat Full Length 1 Answers
  - Chapter 2: Essential Elements of Kaplan Mcat Full Length 1 Answers
  - Chapter 3: Kaplan Mcat Full Length 1 Answers in Everyday Life
  - Chapter 4: Kaplan Mcat Full Length 1 Answers in Specific Contexts
  - Chapter 5: Conclusion
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