

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Identify a conjugate acid-base pair in chemical reactions involving acids and bases.
2. Calculate the pH of a solution from the concentration of H^+ or H_3O^+ ions, and vice versa.
3. Determine the pH of a solution from K_a and K_b .

Vocabulary

acid	1. conjugate base	acid ionization
acid ionization	2. K_a	base ionization
base	3. K_b	pH

Key Equation

1. $pH = -\log[H^+]$ (Equation 19.1)
2. $[H^+] = 10^{-pH}$ (Equation 19.2)

Part A Completion

Write the chemical formula or identify the substance in each space, and then name the substance. Write the chemical formula in the space provided after the name, identifying its nature.

1. A weak acid, $K_a = 1.0 \times 10^{-4}$. _____
2. A weak base, $K_b = 1.0 \times 10^{-4}$. _____
3. A strong acid, $K_a = 1.0 \times 10^{-1}$. _____
4. A strong base, $K_b = 1.0 \times 10^{-1}$. _____
5. A weak acid, $K_a = 1.0 \times 10^{-4}$. _____
6. A weak base, $K_b = 1.0 \times 10^{-4}$. _____
7. A strong acid, $K_a = 1.0 \times 10^{-1}$. _____
8. A strong base, $K_b = 1.0 \times 10^{-1}$. _____
9. A weak acid, $K_a = 1.0 \times 10^{-4}$. _____
10. A weak base, $K_b = 1.0 \times 10^{-4}$. _____
11. A strong acid, $K_a = 1.0 \times 10^{-1}$. _____
12. A strong base, $K_b = 1.0 \times 10^{-1}$. _____
13. A weak acid, $K_a = 1.0 \times 10^{-4}$. _____
14. A weak base, $K_b = 1.0 \times 10^{-4}$. _____
15. A strong acid, $K_a = 1.0 \times 10^{-1}$. _____
16. A strong base, $K_b = 1.0 \times 10^{-1}$. _____

Hydrogen Ions And Acidity Section Review Answers

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Sun**

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