Hess's Law Practice Problems

Please check back for the answers. I will repost this document with the answers at a later date

1. Given the following equations and ΔH^* values, determine the heat of reaction (kJ) at 298 K for the reaction:

$$B_2H_4(g) + 6 Cl_2(g) \rightarrow 2 BCl_3(g) + 6 HCl(g)$$

 $BC1_1(g) + 3 H_2O(1) \longrightarrow H_2BO_1(g) + 3 HC1(g)$ $B_1H_1(g) + 6 H_2O(1) \longrightarrow 2 H_2BO_1(g) + 6 H_2(g)$ $1/2 H_2(g) + 1/2 C1_2(g) \longrightarrow BC1(g)$ $\Delta H^{o}/kJ = -112.5$ $\Delta H^{o}/kJ = -493.4$ $\Delta H^{o}/kJ = -92.3$

Answer: -1376 kJ

2. Given the following equations and ΔH^* values, determine the heat of reaction (kJ) at 298 K for the reaction:

$$2 \text{ OF}_2(g) + 2 \text{ S(s)} \rightarrow \text{SO}_2(g) + \text{SF}_4(g)$$

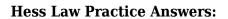
 $OF_2(g) + H_2O(1) \longrightarrow O_2(g) + 2 HF(g)$ $SF_4(g) + 2 H_2O(1) \longrightarrow 4 HF(g) + SO_2(g)$ $S(s) + O_2(g) \longrightarrow SO_2(g)$ $\Delta H^{+}/kJ = -276.6$ $\Delta H^{+}/kJ = -827.5$ $\Delta H^{+}/kJ = -296.9$

Answer: -319.5 kJ

3. Determine $\triangle H^*/kJ$ for the following reaction using the listed enthalpies of reaction:

Hess Law Practice Answers

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