

3.  $\text{S}_8 + 12 \text{O}_2 \rightarrow 8 \text{SO}_3$
4.  $2 \text{N}_2 + \text{O}_2 \rightarrow 2 \text{N}_2\text{O}$
5.  $2 \text{HgO} \rightarrow 2 \text{Hg} + \text{O}_2$
6.  $6 \text{CO}_2 + 6 \text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + 6 \text{O}_2$
7.  $\text{Zn} + 2 \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
8.  $\text{SiCl}_4 + 4 \text{H}_2\text{O} \rightarrow \text{H}_4\text{SiO}_4 + 4 \text{HCl}$
9.  $2 \text{Na} + 2 \text{H}_2\text{O} \rightarrow 2 \text{NaOH} + \text{H}_2$
10.  $2 \text{H}_3\text{PO}_4 \rightarrow \text{H}_4\text{P}_2\text{O}_7 + \text{H}_2\text{O}$
11.  $\text{C}_{10}\text{H}_{16} + 8 \text{Cl}_2 \rightarrow 10 \text{C} + 16 \text{HCl}$
12.  $\text{CO}_2 + 2 \text{NH}_3 \rightarrow \text{OC}(\text{NH}_2)_2 + \text{H}_2\text{O}$
13.  $4 \text{Si}_2\text{H}_3 + 17 \text{O}_2 \rightarrow 8 \text{SiO}_2 + 6 \text{H}_2\text{O}_3$
14.  $2 \text{Al}(\text{OH})_3 + 3 \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 6 \text{H}_2\text{O}$
15.  $4 \text{Fe} + 3 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3$
16.  $\text{Fe}_2(\text{SO}_4)_3 + 6 \text{KOH} \rightarrow 3 \text{K}_2\text{SO}_4 + 2 \text{Fe}(\text{OH})_3$
17.  $2 \text{C}_7\text{H}_6\text{O}_2 + 15 \text{O}_2 \rightarrow 14 \text{CO}_2 + 6 \text{H}_2\text{O}$
18.  $\text{H}_2\text{SO}_4 + 8 \text{HI} \rightarrow \text{H}_2\text{S} + 4 \text{I}_2 + 4 \text{H}_2\text{O}$
19.  $4 \text{FeS}_2 + 11 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$
20.  $2 \text{Al} + 3 \text{FeO} \rightarrow \text{Al}_2\text{O}_3 + 3 \text{Fe}$
21.  $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow 2 \text{Fe} + 3 \text{H}_2\text{O}$
22.  $\text{Na}_2\text{CO}_3 + 2 \text{HCl} \rightarrow 2 \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$

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**Padhraic Smyth**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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