

Advance Study Assignment: Identification of a Compound by Mass Relationships

1. A student attempted to identify an unknown compound by the method used in this experiment. She found that when she heated a sample weighing 0.5032 g, the mass dropped to 0.2663 g. When the product was converted to a chloride, the mass went back up, to 0.3747 g.

- a. Is the sample a carbonate?

Yes ☒ (Circle one)

Please provide your reasoning below.

It is observed that for bicarbonates, a negative while for carbonates $\Delta =$ positive for unknown a negative. This shows unknown to be a bicarbonate.

- b. What are the two compounds that might be in the unknown?



- c. Write the balanced chemical equation for the overall reaction that occurs when each of these two original compounds is converted to a chloride. If the compound is a hydrogen carbonate, use the sum of Reactions 1 and 2. If the sample is a carbonate, use Reaction 2. Write the equation for a sodium salt and then for a potassium salt.



- d. How many moles of the chloride salt would be produced from one mole of the original compound?

1 mole

- e. How many grams of the chloride salt would be produced from one molar mass of the original compound?

Molar masses: NaHCO_3 , 84.007 g/mol Na_2CO_3 , 105.988 g/mol NaCl , 58.44 g/mol

KHCO_3 , 100.118 g/mol K_2CO_3 , 138.21 g/mol KCl , 74.55 g/mol

If a sodium salt, 84.007 g original compound $\rightarrow$ 58.44 g chloride

If a potassium salt, 100.118 g original compound $\rightarrow$ 74.5513 g chloride

(continued on following page)

Experiment 5 Advance Study Assignment

M Mosston



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