

STOICHIOMETRY

Review Questions:

- 1) How many H atoms are there in 45.0 g $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$?
Given: $\text{MM}_\text{N} = 14.01$; $\text{MM}_\text{H} = 1.008$; $\text{MM}_\text{Cr} = 52.00$; $\text{MM}_\text{O} = 16.00$
A) 2.15×10^{23} atoms
B) 6.02×10^{23} atoms
C) 7.52×10^{23} atoms
D) 8.60×10^{23} atoms
- 2) How many moles of calcium are required to produce 5.50 g of calcium phosphate?
The molar mass of calcium phosphate is 310.18 g/mol
A) 1.77×10^{-2} mol
B) 5.32×10^{-2} mol
C) 6.28×10^{-2} mol
D) 4.07×10^{-2} mol
- 3) Combustion analysis of a 12.0-mg sample of ibuprofen produced 9.50 mg of water.
What is the percent hydrogen in ibuprofen?
A) 8.87%
B) 17.7%
C) 4.44%
D) 79.2%

Answers to the above questions: 1 d); 2 b); 3 a)

CHEMISTRY'S ACCOUNTING

Reaction stoichiometry allows us to use **balanced equations** to convert/determine the following:

Mass $\leftrightarrow$ Moles $\leftrightarrow$ Molecules $\leftrightarrow$ Atoms

What do balanced equations tell you?

e.g., $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

- molecules?
- moles?
- masses?
- atoms?

Experiment 3 stoichiometry Answers

JG Myers



Experiment 3stoichiometry Answers:

Embark on a breathtaking journey through nature and adventure with Explore with is mesmerizing ebook, **Experiment 3stoichiometry Answers** . This immersive experience, available for download in a PDF format (*), transports you to the heart of natural marvels and thrilling escapades. Download now and let the adventure begin!

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