

EXPERIMENT 19

ACIDIC AND BASIC SALTS

INTRODUCTION

Water makes up about 60% of the weight of a human body; other animals contain a similar proportion of water, and some plants are almost 95% water. Water covers about 70% of the earth's surface; also, it is in the soil and is present as vapor in the atmosphere. Consequently, reactions that occur in aqueous solutions and reactions in which water participates are very important to us; many are essential to the life and growth of organisms. These reactions might affect the pH of water by using or producing hydronium or hydroxide ions.

It is apparent from the formulas of some compounds that they will produce H_3O^+ or OH^- in aqueous solution. An Arrhenius acid contains hydrogen, which is usually written first in the formula to emphasize its ability to react with water producing H_3O^+ . An ionic Arrhenius base contains hydroxide ions, which separate from the positive ions when the base dissolves in water.

Salts are ionic compounds that contain a cation from a base and an anion from an acid. Even though most salts do not contain H or OH, many of them do react with water changing its pH. Such a reaction is called the *hydrolysis of a salt*, the reaction of an ion of a dissolved salt with water producing H_3O^+ or OH^- .

The hydrolysis of a negative ion produces the negative ion of water. This reaction occurs if the anion is the conjugate base of a weak acid. For example, HCN is a weak acid; $K_a = 4.9 \times 10^{-10}$:



The acid molecule HCN has little tendency to donate protons; its conjugate base, CN^- , is a good proton acceptor. Consequently, when a salt containing CN^- , such as NaCN, dissolves in water, a Brønsted-Lowry acid-base reaction occurs between H_2O (acid) and CN^- (base); that is, a proton is transferred from H_2O to CN^- :



The resulting solution is basic, because hydroxide ions in excess of those produced by the autoionization of water are present.

The hydrolysis of a positive ion produces the positive ion of water. This reaction occurs if the cation is the conjugate acid of a weak base. Molecular bases such as NH_3 are weak bases:



For this reaction, $K_b = 1.8 \times 10^{-5}$. The ammonium ion can donate H^+ to a chemical species that is capable of acting as a base. Consequently, when a salt containing NH_4^+ dissolves in water, NH_4^+ and H_2O react in a Brønsted-Lowry fashion:



Some hydrated metal ions also hydrolyze producing acidic solutions. The strong attraction between a small, highly charged, metal ion and the oxygen atom of a bound water molecule weakens an oxygen-hydrogen bond in the bound water molecule. As a result, the hydrate acts as a proton donor:

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John Bernard Ekeley



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Laboratory Experiments in Chemistry Newton Henry Black,1921 *A Laboratory Study of Household Chemistry* Mary Ethel Jones,1921 **Laboratory Experiments in Chemistry to Accompany Black and Conant's "Practical Chemistry"** Newton Henry Black,1920 **Experiments and Exercises in Basic Chemistry** Steven Murov,Brian Stedjee,2000-02-03

Laboratory Manual for Principles of General Chemistry Jo Allan Beran,2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures *Soil Reaction in Relation to Calcium Adsorption* Charles Oscar Swanson,1922 *An Elementary Experimental Chemistry* John Bernard Ekeley,1900 **Secondary School Science Teaching Practices** Horatio Seymour Fowler,1964 **Practical Physical Chemistry** Dr. M. Satish Kumar, Physical Chemistry deals with the relations between the physical properties of substances and their composition The present book is intended to serve as a practical manual for undergraduate and post graduate students I have attempted to assemble the list of experiments from my experience and also have drawn upon the experience of the students who have undergone these laboratory courses and felt the inadequacy of the existing syllabus I am aware that I have not yet exhausted all the experiments that they wanted to place in this book but I had to make a selection keeping the size in consideration This manual is largely structured around the standard experiments of physical chemistry Detailed information on instrumentation kinetics experimental methods and data analysis has been covered I will be happier to take all comments and incorporate them in the further editions Science Olympiad Workbook - Class 7 Swastick Book Box,Preeti Garg,2021-01-07 SBB Olympiad Workbook helps students to prepare for Olympiad exams through meticulously designed contents The workbook helps the child to focus on the basics of the subject and promotes logical brain development that improves child overall logical thought process and problem solving skills Feature List Olympiads Question Bank Preparatory book for class 7 students Prepared based on NCERT CBSE ICSE IGCSE IB and state board syllabus Questions patterned for Science Olympiad Foundation Indian Talent Olympiad SilverZone ASSET EduHeal NSTSE Contains chapter wise MCQs 2 Model test papers Answers Key Tags Science Olympiad Books For Class 7 Olympiad NSO Book For Class 7 Olympiad Science Class 7 Olympiad Question Papers Class 7 Science Olympiad Class 7 Books Olympiad Science Workbook Class 7 Olympiad Excellence Guide Class 7 Science

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