

Name \_\_\_\_\_ Section \_\_\_\_\_

**Experiment 41****Advance Study Assignment: Preparation of Aspirin**

1. Calculate the theoretical yield of aspirin to be obtained in this experiment, starting with 2.0 g of salicylic acid and 5.0 mL of acetic anhydride (density = 1.08 g/mL).

2. If 1.8 g of aspirin were obtained in this experiment, what would be the percentage yield?

3. The name acetic anhydride implies that the compound will react with water to form acetic acid. Write the equation for the reaction.

4. Identify R and R' in Equation 1 when the ester, aspirin, is made from salicylic acid and acetic acid.

5. Write the equation for the reaction by which aspirin decomposes in an aqueous ethanol solution.

# Experiment 41 Preparation Aspirin Answers

**Michal Rosen-Zvi**



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