

Exercise E. Combined Gas Law

Using the combined gas law, $\frac{\text{Pressure}(P) \times \text{Volume}(V)}{\text{Temperature}(T)} = \text{Constant}$

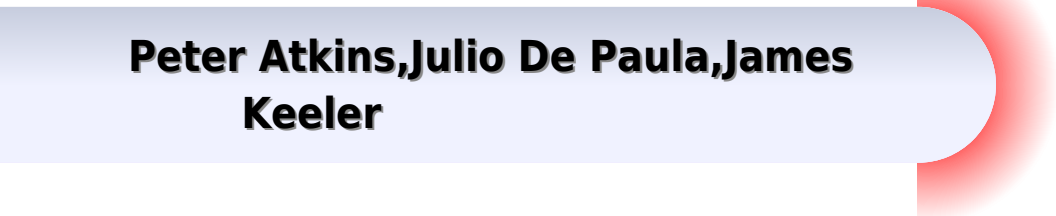
calculate the parameter that is missing in each row of the following table. Enter the appropriate form of the combined gas law in the first column as shown in the table. Be sure to enter the appropriate units for each result and use the appropriate units when calculating the result.

Equation	T1	P1	V1	T2	P2	V2
$\frac{P_2 \times V_2 \times T_1}{P_1 \times V_1} = T_2$	300. K	2.0 atm	2.0 liters	1200 K	4.0 atm	4.0 liters
$\frac{P_1 \times V_1 \times T_2}{P_2 \times T_1} = V_2$	400. K	1.50 atm	2.50 liters	600. K	2.00 atm	2.8 liters
$\frac{P_1 \times V_1 \times T_2}{V_2 \times T_1} = P_2$	500 K	3.00 atm	6.00 liters	400. C	97 atm	250 mL
	400 K	760. mm Hg	1.0 liter		1.5 atm	2.00 liters

A gas that has a volume of 10 L

Exercise 12 Gas Laws Answers

**Peter Atkins,Julio De Paula,James
Keeler**



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