

1. Define half-life. *(1 mark)*
2. How is the half-life of a radioisotope similar to a sporting tournament in which the losing team is eliminated? *(2 marks)*
3. The half-life of radium-226 is 1600 years. If a sample of radium-226 has an original activity of 200 Bq, what will its activity be after 4800 years? *(2 marks)*
4. Sodium-24 has a half-life of 15 hours. If a sample of sodium-24 has an original activity of 500 Bq, what will its activity be after 60 hours? *(2 marks)*
5. After 42 days the activity of a sample of phosphorus-32 has decreased from 400 Bq to 50 Bq. What is the half-life of phosphorus-32? *(2 marks)*
6. The half-life of radon-222 is 3.8 days. What was the original activity if it has an activity of 10 Bq after 7.6 days? *(2 marks)*
7. The half-life of thorium-227 is 19 days. How many days are required for 75% of a sample to decay? *(2 marks)*
8. The half-life of protactinium-234 is 6.75 hours. What percentage of a sample will remain after 27 hours? *(2 marks)*
9. A rock once contained 1.0 mg of uranium-238, but now contains only 0.25 mg. Given that the half-life for uranium-238 is 4.5×10^9 (4.5 billion) years, how old is the rock? *(2 marks)*
10. The half-life of tritium (hydrogen-3) is 12.3 years. If 48.0 mg of tritium is released from a nuclear power plant during the course of a mishap, what mass of the sample will remain after 49.2 years? *(2 marks)*

Total marks: 19

Instructional Fair Inc Physical Science Half Life Calculations

Rachel S Tattersall



Instructional Fair Inc Physical Science Half Life Calculations:

The book delves into Instructional Fair Inc Physical Science Half Life Calculations. Instructional Fair Inc Physical Science Half Life Calculations is a vital topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Instructional Fair Inc Physical Science Half Life Calculations, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Instructional Fair Inc Physical Science Half Life Calculations
 - Chapter 2: Essential Elements of Instructional Fair Inc Physical Science Half Life Calculations
 - Chapter 3: Instructional Fair Inc Physical Science Half Life Calculations in Everyday Life
 - Chapter 4: Instructional Fair Inc Physical Science Half Life Calculations in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, this book will provide an overview of Instructional Fair Inc Physical Science Half Life Calculations. The first chapter will explore what Instructional Fair Inc Physical Science Half Life Calculations is, why Instructional Fair Inc Physical Science Half Life Calculations is vital, and how to effectively learn about Instructional Fair Inc Physical Science Half Life Calculations.
 3. In chapter 2, the author will delve into the foundational concepts of Instructional Fair Inc Physical Science Half Life Calculations. The second chapter will elucidate the essential principles that must be understood to grasp Instructional Fair Inc Physical Science Half Life Calculations in its entirety.
 4. In chapter 3, this book will examine the practical applications of Instructional Fair Inc Physical Science Half Life Calculations in daily life. This chapter will showcase real-world examples of how Instructional Fair Inc Physical Science Half Life Calculations can be effectively utilized in everyday scenarios.
 5. In chapter 4, the author will scrutinize the relevance of Instructional Fair Inc Physical Science Half Life Calculations in specific contexts. The fourth chapter will explore how Instructional Fair Inc Physical Science Half Life Calculations is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Instructional Fair Inc Physical Science Half Life Calculations. This chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Instructional Fair Inc Physical Science Half Life Calculations.

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web answer the following questions based on the solubility curve below which salt is least soluble in water at 20 c 2 how many grams of potassium chloride can be dissolved in 200 g of water at 80 c 3 at 40 c how much potassium nitrate can be dissolved in 300 g of water w 4 which salt shows the least change

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web different substances q why do temperatures range from 0 100 c a water is a liquid at these temperatures solubility curves tell us how much solid will dissolve at a given temperature they also allow certain calculations

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web chemistry solubility curves virtual lab curves to clearly show the solubility of the four ionic compounds in the data table below the following tips will help you make the best possible graph make the graph large it should fill three fourths of the graph paper

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web solvents the larger volume of the mixture and solutes the smaller volume of the mixture for example a hot chocolate is a solution in which the solute the chocolate powder is dissolved in the solvent the milk or water the solute and solvent can be

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web solubility curve lab sheet pre lab 1 define the terms solute solvent and solution in your definition describe a solution found in your everyday life solute a solute is the component of the mixture that is dissolved such as the powdered drink

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web question solubility curve of potassium nitrate purpose in this experiment you will determine the solubilities of potassium nitrate in water at various temperatures and construct a solubility curve for knoy

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web they should be able to come up with the following answers stir the water increase the temperature and increase the amount of water alternately you may challenge the students to a contest

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web use a solubility curve to determine whether a solution is saturated or unsaturated and to estimate the amount of excess solute that remains undissolved in the solution

2 3 solubility lab report chemistry libretexts - Feb 15 2023

web give results of all calculations in blue cells you must put your answers in the indicated cells as the grade book pulls your data from those cells do not put any units in the number cells use the unit boxes given graph tab transfer from data page the

temp and concentration of salt in units of g salt 100 g water

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web answer key 4th edition ap psychology unit 1 study guide 1 3 discussion population samples and bias lab solubility curves of a solid purpose to determine how the solubility of a salt ammonium chloride varies with temperature in order to create a solubility curve based on observed data

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web 30 09 99 solubility curves class data write down the data answer questions on bottom of the sheet results of lab solubility of KClO_3 unsaturated saturation and solubility curves in this experiment we expressed concentration in units of g solute g H_2O putting concentrations in g solute g solution would be tedious and impractical see fig 2

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web use the solubility curve to answer these questions determine how much potassium nitrate will dissolve in 20 g of water at 40 c 200 cm³ of saturated lead ii nitrate solution was prepared at a temperature of 90 c

2 2 solubility lab chemistry libretexts - Aug 21 2023

web collect experimental data and create a solubility curve by the end of this lab students should be able to properly use an analytical balance to measure mass set up an experimental work station to measure the solubility of a salt in water as a function of the temperature generate a workbook using google sheets prior knowledge

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web sample answer sugar has a much greater solubility in hot water than cold water so its solubility curve probably resembles the sloping solubility curve of potassium nitrate temperature solubility g 100 ml 10 c 36 g 100 ml 20 c 36 g 100 ml 30 c 37 g 100 ml 40 c 37 g 100 ml 50 c 38 g 100 ml 60 c 38 g 100 ml 70 c 39 g 100 ml 80

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web 2 days ago study with quizlet and memorize flashcards containing terms like to test the hypothesis above you changed the of the solvent and measured the of a solid therefore the independent variable was and the dependent variable was the results of this lab experiment the hypothesis and more

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web may 29 2018 figure 2 the graph shows the solubility curve of potassium nitrate with corresponding temperatures as the temperature increases so does the solubility of the solution a y 0 2 1 61 b y 0 2 1 61 0 60 2 1 60 61 0 40 2 1 40 61 8 5 g 100 ml 4 6 g 100 ml

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web mar 15 2021 this experiment will further develop your understanding of thermodynamics while relating it to two

concepts you already have studied equilibrium and solubility you will determine three thermodynamic values associated with the solubility of potassium nitrate the reaction we will be studying is the dissolution of potassium nitrate in water

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web analysis construct a solubility curve using the mass of solute that will dissolve in 100 g of water at a specified temperature in this lab 5 ml of water was used assume 10 g ml of water take the collected data make it proportional to get the solubility of the KNO_3 graph temperature vs solubility using a best fit curve

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web part i construct a single graph with multiple solubility curves to clearly show the solubility of the four ionic compounds in the data table below the following tips will help you make the best possible graph