#### Chapter 2: Atoms, Molecules and Ions

Dalton's Atomic Theory (1805)

- All matter consists of atoms; tiny indivisible particles of an element that cannot be created or destroyed.
- 2. Atoms of one element cannot be subdivided, created, or destroyed during chemical reactions (ipmortion)
- 3. Atoms of an element are identical in mass and other properties and are different from the atoms of any other element. (wnonw)
- 4. Compounds result from the chemical combination of a specific ratio of atoms of different elements.

Atomic Structure: Flectrons

- J.J. Thomson (1897) Cathode Ray Tube Experiment
- Observations:
  - Ray bends in magnetic field
  - Ray bends toward positive plate in electric field
  - Ray is identical for any cathode
- Conclusions:
  - Ray consists of charged particles
  - Ray consists of negative particles
  - These particles are found in ALL matter

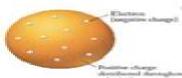
#### Cathode Ray Tube experiment using magnets led to the discovery of electrons

Thomson's Model of the Atom

- Plum-pudding model:
  - E- distributed throughout diffuse, positively charged "squishy" sphere

Atomic Structure: The Nucleus

- Rutherford's Experiment
  - Bombard a thin gold foil with positively charged particles to test Thomson's model of the atom.
  - Observations:
    - Atom is made up of smaller subatomic particles
      - Protons
      - Neutrons
      - Electrons (mass of only ab 1/2000th of other particles)
  - Conclusions



# **Genral Chemistry Study Guide**

**Jack Rudman** 

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