Name: ______ Date: _____ Period: ____

Ch 11 Worksheet 5 - Thermochemical Stoichiometry

Given the equation: 2 H₂ (g) + O₂ (g) → 2 H₂O (I)

 $\Delta H = -571.7 \text{ kJ}$

- a. How much energy is produced if 100.0 grams of H₂ are converted to water?
- b. How many grams of H₂ would be required to produce 5000.0 kJ?
- Given the equation: N₂(g) + 3H₂(g) → 2NH₃(g)

 $\Delta H_{\rm EXR} = -92.22 kJ$

- a. Is the formation of ammonia endothermic or exothermic?
- b. What is the enthalpy change associated with the formation of 1.0 mol of NH₃?
- c. Is the decomposition of ammonia endothermic or exothermic?
- d. What is the enthalpy change associated with the decomposition of 4.00 mol NH₃?
- e. How much energy will be released when 12.0 g H₂(g) react with excess N₂(g) to form NH₃(g)?
- Given the equation: 2Ca(s) + O₂(g) → 2CaO(s)

 $\Delta H = -635.09 kJ$

- a. How much energy is released during the production of 5.0 g of CaO?
- b. How much energy must be added to CaO to produce 64.0 g of O₂(g)?
- Given the equation: 4Fe(s) + 3 O₂(g) → 2Fe₂O₂(s)

 $\Delta H = -824.2 \text{ kJ}$

- a. Assuming that the iron is converted to iron (III) oxide as in the above equation, how much heat can be provided by a hot pack that contains 50.0 g of iron?
- b. If the hot pack from part a, gives off heat at the rate of 42.0 kJ /hour, how long will it stay warm?

Answers:

1a. 14,150 kJ produced 1b. 35.3g 2a. 2b. -46kJ 2c. 2d. 184 kJ 2e. 183 kJ 3a. 28 kJ released

3b. 1270 kJ 4a. 185 kJ provided 4b. 4.40 hours

Energy Ws 5 Thermochemical Equations And Stoichiometry

Akash Kumar Bhoi, Karma Sonam Sherpa, Akhtar Kalam, Gyoo-Soo Chae

Energy Ws 5 Thermochemical Equations And Stoichiometry:

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