

# Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$\text{C}_{40.0} \text{H}_{6.67} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 3.33 \text{ moles C}$$

$$\text{H}_{6.67} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 6.62 \text{ moles H}$$

C<sub>3</sub>H<sub>5</sub>

2. A compound is found to contain 50.0 g of potassium, 43.8 g of oxygen, and 6.7 g of hydrogen. Find the empirical formula.

$$\text{K}_{50.0} \text{O}_{43.8} \text{H}_{6.7} \times \frac{1 \text{ mole}}{39.10 \text{ g K}} = 1.28 \text{ moles K}$$

$$\text{O}_{43.8} \times \frac{1 \text{ mole}}{16.00 \text{ g O}} = 2.74 \text{ moles O}$$



$$\text{H}_{6.7} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 6.64 \text{ moles H} \quad \text{K}_2 \text{O}_4 \text{H}_1$$

3. A 10.0 g sample of a compound, a possible alkane, is found to contain 8.56 g of carbon, 1.44 g of hydrogen, and 0.00 g of oxygen. Find the empirical formula.

$$\text{C}_{8.56} \text{H}_{1.44} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 0.713 \text{ moles C}$$



$$\text{H}_{1.44} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 1.43 \text{ moles H} \quad \text{C}_{0.713} \text{H}_{0.119} \times 7.17 = \text{C}_5 \text{H}_1$$

$$\text{C}_5 \text{H}_1 \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 5.74 \text{ moles C} \quad \text{C}_{5.74} \text{H}_1 \times 0.174 = \text{C}_1 \text{H}_0.1$$

4. A 25.00 g sample of an unknown compound contains 15.38 g carbon, 4.08 g hydrogen, and 5.54 g oxygen. Find the empirical formula.

$$\text{C}_{15.38} \text{H}_{4.08} \text{O}_{5.54} \times \frac{1 \text{ mole}}{12.01 \text{ g C}} = 1.28 \text{ moles C}$$



$$\text{H}_{4.08} \times \frac{1 \text{ mole}}{1.008 \text{ g H}} = 4.04 \text{ moles H}$$

$$\text{O}_{5.54} \times \frac{1 \text{ mole}}{16.00 \text{ g O}} = 0.346 \text{ moles O} \quad \text{C}_{1.28} \text{H}_{4.04} \text{O}_{0.346} \times 2.89 = \text{C}_3 \text{H}_{11.6} \text{O}_1$$

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