

## EMPIRICAL AND MOLECULAR FORMULA

- 1) What is the empirical formula of a compound that contains 46.3 % lithium and 53.7% oxygen?
- 2) What is the empirical formula of a compound that contains 15.9 % boron and 84.1 % fluorine?
- 3) Phosphorus reacts with oxygen to give a compound that is 43.7% phosphorus and 56.4% oxygen. What is the empirical formula of the compound?
- 4) An inorganic salt is composed of 17.6% sodium, 39.7% chromium and 42.8% oxygen. What is the empirical formula of this salt?
- 5) Compound X contains 69.9% carbon, 6.86% hydrogen and 23.3% oxygen. Determine the empirical formula of X.
- 6) Oxalic acid has the empirical formula  $\text{CHO}_2$ . Its molar mass is 90 g/mol. What is the molecular formula of oxalic acid?
- 7) The empirical formula of codeine is  $\text{C}_{18}\text{H}_{21}\text{NO}_3$ . If the molar mass of codeine is 299 g/mol, what is its molecular formula?
- 8) A compound's molar mass is 240.28 g/mol. Its percentage composition is 75.0% carbon, 5.05% hydrogen and 20.0% oxygen. What is the compound's molecular formula?
- 9) The wintergreen plant produces methyl salicylate, or oil of wintergreen. It can also be prepared easily in the laboratory. Methyl salicylate is 63.1% carbon, 5.31% hydrogen and 31.6% oxygen. Calculate the empirical formula of this compound.
- 10) An inorganic salt is made up of 38.8% calcium, 20.0% phosphorus, and 41.2% oxygen.
  - a) What is the empirical formula of the compound?
  - b) On further analysis, each formula unit of this salt is found to contain two phosphate ions. Predict the molecular formula of this salt.

### Answers

- 1)  $\text{Li}_2\text{O}$
- 2)  $\text{BF}_3$
- 3)  $\text{P}_2\text{O}_5$
- 4)  $\text{Na}_2\text{Cr}_2\text{O}_7$
- 5)  $\text{C}_{12}\text{H}_{14}\text{O}_3$
- 6)  $\text{C}_2\text{H}_2\text{O}_4$
- 7)  $\text{C}_{18}\text{H}_{21}\text{NO}_3$
- 8)  $\text{C}_{15}\text{H}_{12}\text{O}_3$
- 9)  $\text{C}_8\text{H}_8\text{O}_3$
- 10) a)  $\text{Ca}_3\text{P}_2\text{O}_8$  b)  $\text{Ca}_3(\text{PO}_4)_2$

# Empirical And Molecular Formula Problems

**Thomas Griffiths**



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