

*Lab (25 pts):
Empirical Formula
of Magnesium Oxide*

Name _____
Partner _____
Period _____ Date _____

Purpose: To experimentally determine the empirical formula for magnesium oxide, Mg_xO_y .

Introduction:

By now you should know that a **molecular formula** is a group of elemental symbols and subscript numbers, which represent the composition of a molecule. The molecular formula shows you how many of each atom can be found in a certain molecule. Some examples of molecular formulas, that you are most likely familiar with, are shown in Table A.

Table A - Some Molecular Formulas and the Information they Convey	
Formula	Composition
H_2O	Total of 3 atoms in each molecule - 1 oxygen and 2 hydrogen
O_2	Total of 2 atoms in each molecule - 2 oxygen atoms
CO_2	Total of 3 atoms in each molecule - 1 carbon and 2 oxygen
$\text{C}_6\text{H}_{12}\text{O}_6$	Total of 24 atoms in each molecule - 6 carbon, 12 hydrogen, and 6 oxygen.

While the molecular formula shows the molecule as it actually exists, the **empirical formula** for a compound shows the simplest whole number ratio for the elements in the compounds, like reducing a fraction. To find the empirical formula for a compound, find the largest whole number that can be divided into each of the subscripts evenly (greatest common factor). If no number greater than 1 will go into any of the subscripts evenly, then the empirical formula and the molecular formula for the compound are the same. Table B shows examples of some empirical and molecular formulas.

Table B - Comparing Empirical and Molecular Formulas			
Compound	Molecular Formula	GCF	Empirical Formula
Water	H_2O	1	H_2O
Hydrogen Peroxide	H_2O_2	2	HO
Glucose	$\text{C}_6\text{H}_{12}\text{O}_6$	6	CH_2O
Methane	CH_4	1	CH_4
Ethane	C_2H_6	2	CH_3
Octane	C_8H_{18}	2	C_4H_9

In this lab, you will perform the following reaction:



where you will determine the subscripts, x and y . The mass of the solid (magnesium) will increase as the oxygen gas combines with magnesium metal to form magnesium oxide. You will measure the mass of Mg before the reaction and the mass of Mg_xO_y after the reaction to determine the mass of O that reacted from the air, and the mass of Mg and O in magnesium oxide. Then you will calculate magnesium oxide's empirical formula.

Procedure: Goggles are **ABSOLUTELY** required!! Don't touch a **HOT** crucible--it will burn you!

- 1) Obtain a stainless steel crucible and lid, and record the mass of the crucible and lid together (watch units & sig figs). Slight discoloration of the crucible is normal and will not affect your results. We are using special balanced for this lab, which measure mass to 0.001 g. *Record all masses to 0.001 g (3 decimal pts).* Use gloves or a crucible tong to handle the crucible. When carrying the crucible hold it firmly with crucible tongs, and hold a piece of metal gauze beneath the crucible in case it slips.



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Jo Allan Beran



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