

Lab: Single and Double Replacement Reactions

Purpose: To predict products and observe various single and double replacement reactions.

Hypotheses: Single Replacement Reactions: Use the **activity series** (p.268 - Honens) and your understanding of single replacement reactions to predict products for the following reactions. If no reaction will occur, then write "no rxn" on the product side of the equation. Remember to **BALANCE** the equation afterwards.

- 1) $\text{HCl} + \text{Mg} \rightarrow$ _____
- 2) $\text{HCl} + \text{Cu} \rightarrow$ _____ (hint: copper could be +2 in this compound)
- 3) $\text{HCl} + \text{Fe} \rightarrow$ _____ (hint: iron will be +3 in this compound)
- 4) $\text{HCl} + \text{Zn} \rightarrow$ _____

Hypotheses: Double Replacement Reactions: Use the **solubility chart** and your understanding of double replacement reactions to predict products for the following reactions (see Table 2 on the back of this paper...we're only fully writing out equations for 4 of the 16 you will test). If both of the possible products are aqueous, then write "no rxn" on the product side of the equation. Otherwise, clearly state which product will be aqueous and which will form a solid precipitate. Remember to **BALANCE** the equation afterwards.

- 1) $\text{Na}_2\text{CO}_3(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow$ _____
- 6) $\text{NaOH}(\text{aq}) + \text{Fe}(\text{NO}_3)_3(\text{aq}) \rightarrow$ _____
- 11) $\text{Na}_3\text{PO}_4(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$ _____
- 16) $\text{MgSO}_4(\text{aq}) + \text{Zn}(\text{NO}_3)_2(\text{aq}) \rightarrow$ _____

Procedures:

1. For each single replacement reaction, place a sample of each metal in your well plate and then place 5-6 drops of HCl (aq) on top of it. Record your observations.
2. For each double replacement reaction, place two drops of each solution carefully on the transparency, covering each "X" appropriately over your "Super High-Tech Patented Reaction Surface." Record your observations.
3. If a precipitate formed, clearly write the formula for the precipitate on the line provided.
4. Clean up your workstation according to your instructor's directions and answer all questions.

Observations/Data:

Table 1: Single Replacement Reaction Observations

	HCl (aq)
Mg	1 Observations:
Cu	2 Observations:
Fe	3 Observations:
Zn	4 Observations:

Lab 27 Double Replacement Reactions

Jennifer Louise Pakkala



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Table of Contents Lab 27 Double Replacement Reactions

1. Understanding the eBook Lab 27 Double Replacement Reactions
 - The Rise of Digital Reading Lab 27 Double Replacement Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab 27 Double Replacement Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab 27 Double Replacement Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab 27 Double Replacement Reactions
 - Personalized Recommendations
 - Lab 27 Double Replacement Reactions User Reviews and Ratings
 - Lab 27 Double Replacement Reactions and Bestseller Lists
5. Accessing Lab 27 Double Replacement Reactions Free and Paid eBooks

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 - Lab 27 Double Replacement Reactions Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab 27 Double Replacement Reactions
 - Highlighting and Note-Taking Lab 27 Double Replacement Reactions
 - Interactive Elements Lab 27 Double Replacement Reactions
 8. Staying Engaged with Lab 27 Double Replacement Reactions
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab 27 Double Replacement Reactions
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab 27 Double Replacement Reactions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Lab 27 Double Replacement Reactions
 - Setting Reading Goals Lab 27 Double Replacement Reactions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Lab 27 Double Replacement Reactions
 - Fact-Checking eBook Content of Lab 27 Double Replacement Reactions
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Interactive and Gamified eBooks

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