ACID-BASE RELATIVE STRENGTHS in Aqueous Solutions at 25° C

Name of Acid	me of Acid Acid		Base		Ka	
Hydriodic	1	HI	\rightarrow	H* + I-		>> 1
Hydrobromic	181	HBr	\rightarrow	H* + Br		>> 1
Hydrochloric	Strongest Acid	HCI	\rightarrow	H+ + CI-		>> 1
Nitric	臺	HNO,	->	H+ + NO.		> 1
Sulfuric	-26	H ₂ SO ₄	\rightarrow	H+ + HSO4		> 1
Oxalic		H ₂ C ₂ O ₄	=	H+ + HC2O4		5.90 x 10 ⁻²
Sulfurous		H,SO,	===	H+ + HSO3-	3	1.54 x 10 ⁻²
Hydrogen				-	Weakest Base	
Sulfate Ion		HSO4	-	H+ + SO42-	3	1.20 x 10 ⁻²
Phosphoric		H ₃ PO ₄	-	H+ + H,PO4-	~	7.52 x 10 ⁻³
Nitrous		HNO,	\Rightarrow	H+ + NO2-		4.6 x 10 ⁻⁴
Hydrofluoric		HF	-	H+ + F-		3.53 x 10 ⁻⁴
Formic	32	HCO ₂ H	===	H+ + HCO2		1.77 × 10-4
Benzoic	Weak Acid	C,H,CO,H	-	H+ + C.H.CO.		6.46 x 10 ⁻⁵
Acetic	W.	CH ₃ CO ₂ H		H+ + CH3CO2	Total Control	1.76 × 10 ⁻⁵
Carbonic Hydrogen		H ₂ CO ₃	-	H+ + HCO3-	Weak Base	4.30 x 10 ⁻⁷
Sulfite Ion		HSO ₃	\Rightarrow	H+ + SO32-		1.02 × 10 ⁻⁷
Dihydrogen						
Phosphate Ion		H ₂ PO ₄	=	H* + HPO42-		6.23 x 10 ⁻⁸
Boric		H ₃ BO ₃	-	H+ + H2BO3-		7.3 x 10 ⁻¹⁰
Ammonium Ion		NH4"	\Rightarrow	H. + NH".		5.64 × 10 ⁻¹⁰
Hydrogen					- 31	
Carbonate Ion	38	HCO,	===	H* + CO32-	if Bar	5.61 x 10 ⁻¹¹
Monohydrogen	¥				8	
Phosphate Ion	Weakest Acid	HPO ₄ 2-	-	H+ + PO43-	3	2.2 × 10 ⁻¹³
Water	-	H ₂ O	-	H+ + OH-		1.0 × 10 ⁻¹⁴

For space considerations, all reactions omit water as a reactant HA + H,O == H,O* + A*



Lab 16 Creating Acid Base Strength Table

Sabine Zange

Lab 16 Creating Acid Base Strength 7	Table:
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