

1. An unknown compound was found to have a percent composition as follows:
47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula?

$$\begin{aligned} K &= 47.0 \left(\frac{1}{39} \right) = 1.205 / 1.2 = 1 \\ C &= 14.5 \left(\frac{1}{12} \right) = 1.208 / 1.2 = 1 \\ O &= 38.5 \left(\frac{1}{16} \right) = 2.406 / 1.2 = 2 \end{aligned} \quad \boxed{KCO_2}$$

2. A sample of iodine chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the iodine compound?

$$\begin{aligned} I &= .2596 \left(\frac{1}{127} \right) = .00204 / .0012 = 1 \\ Cl &= .2404 \left(\frac{1}{35.5} \right) = .0067 / .0012 = 3 \end{aligned} \quad \boxed{ICl_3}$$

3. Nitrogen and oxygen form an extensive series of oxides with the general formula N_xO_y . One of them is a blue solid that comes apart, reversibly, in the gas phase. It contains 36.84% N. What is the empirical formula of this oxide?

$$\begin{aligned} N &= 36.84 \left(\frac{1}{14} \right) = 2.631 / 2.63 = 1 + 2 = 2 \\ O &= 63.16 \left(\frac{1}{16} \right) = 3.9475 / 2.63 = 1.5 + 2 = 3 \end{aligned} \quad \boxed{N_2O_3}$$

4. Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol?

$$\begin{aligned} C &= 60 \left(\frac{1}{12} \right) = 5 \text{ Atoms} = 5 \\ H &= 13.4 \left(\frac{1}{1} \right) = 13.4 / 1.4 = 10 \\ O &= 26.6 \left(\frac{1}{16} \right) = 1.66 / 1.6 = 1 \end{aligned} \quad \boxed{C_5H_{10}O}$$

5. What is the empirical formula of a compound that contains 0.783g of Carbon, 0.186g of Hydrogen and 0.521g of Oxygen?

$$\begin{aligned} C &= .783 \left(\frac{1}{12} \right) = .06525 = 2 \\ H &= .186 \left(\frac{1}{1} \right) = .186 = 6 \\ O &= .521 \left(\frac{1}{16} \right) = .03256 = 1 \end{aligned} \quad \boxed{C_2H_6O}$$

6. What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\begin{aligned} Au &= 89.14 \left(\frac{1}{197} \right) = .452 / .3 = 1 + 2 = 2 \\ O &= 10.8 \left(\frac{1}{16} \right) = .675 / .3 = 1.5 + 2 = 3 \end{aligned} \quad \boxed{Au_2O_3}$$

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