Lithium bromide is an ionic compound. It can be electrolysed when it is molten or in aqueous solution. It cannot be electrolysed as a solid.

- (a) Solid lithium bromide is a poor conductor of electricity. The ions cannot move to the electrodes, they are held in an ionic lattice by strong forces.
 - (i) Describe the motion of the ions in the solid state.

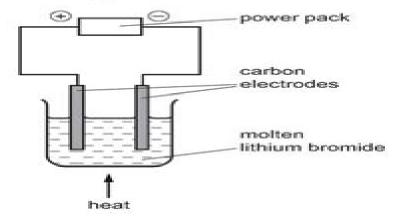
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(ii) Define the term ionic bonding.

(iii) What is meant by the term ionic lattice?

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[2]
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(b) The diagram shows the electrolysis of molten lithium bromide.



- (i) Mark on the diagram the direction of the electron flow.
- (ii) Write an ionic equation for the reaction at the negative electrode (cathode).

[1]

- (iii) Write an ionic equation for the reaction at the positive electrode (anode).
-[2]
- (iv) Which ion is oxidised? Explain your answer.

......[2]

Igcse Year Science Test Papers

David Watson, Helen Williams

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