#### CHEMISTRY I EQUATION SHEET

	- 5	olubility Rules		
Negative Ion F	dus Po	sitive Ion	Form a Compound which is	
		la ions Soluble		
		nium ion Soluble		
Nitrate + an	y positiv	ve ion Soluble		
Acetate		Ag or Hg2"  Ag ",Ptb2",Hg" or Cu"	Soluble	
		Ag Pb Hg or Cu	Not Se	dutole
bromide or iodide + An	y other	positive ion Soluble		
Sulfate		Ca <sup>2*</sup> ,Sr <sup>2*</sup> ,Ba <sup>2*</sup> , Ra <sup>2*</sup> ,Ag or Pb <sup>2*</sup>	Not	Solut
	+ A	ny other positive ion Solubi	6	
		group 1a ions or ammonium Soluble		
Sulfide		Be Mg Ca Srl Ba or Ra	Soluble	
	Any other positive ion Not Soluble			
Hydroxide	+	group 1a ions or ammonium Soluble		
	+ Any other positive ion Not Soluble			
Carbonate, Chromate,		group 1a ions or ammonium Soluble		
Phosphate, or Sulfite + An	y other	positive ion Not Soluble		

Equations	
D = m/v Molarity = mol V P, V, = P, V, T, T, Q = mc ΔT nRTPV - F = I (1/2) * n = Total Time Half Life Time	c = specific heat d = density I = initial amount F = final amount m = mass M = motarity n = number of moles & number of half lives P = pressure R = ideal gas constant Q = heat energy T = temperature V = volume X = amount of matter left

	Constants	
	1 atm = 10	3.7 kPa Pressur
	1 atm = 760 mm Hg	
	R = 8.314 L x	kPa / mol*C Uni
Constant or 0.0	821 L x atm / moi*0	G.
Avogadro's number	6.02	×
Specific Heats	Water =	4,180
(J/g^C)	Aluminum = 0.895	
	Copper = 0.387	
	Iron = 0.448	

Au	otivity
	ories in
	ecreasing
Aus	otivity
	Metals
	thium
	ubidium
	otassium -
	alicium
Sk	odium
NA.	agnesium
	uminum
BA6	anganese
Z	me
line	om:
	cket
The	
Le	haid
H-Ry	ydrogen
Co	оррег
54	Ivver
<b>CFRA</b>	etinum
GH.	olid
	inmetals
	Commiscion
C	hilorine
Br	comine
icos	dine

HEET	
Polyatomic	lons
Name Formula	
Ammonium	NH. "1
Acetate	C <sub>2</sub> H <sub>2</sub> O <sub>2</sub> <sup>-1</sup>
Carbonate	CO3 -2
Bicarbonate or Hydrogen carbonate	HCO <sub>3</sub> ·1
Hypochiorite Chioride.	CIO 1
Chilorite	CIO <sub>2</sub> -1
Chlorate	CIO <sub>3</sub> -1
Perchlorate	CIO <sub>4</sub> <sup>-1</sup>
Hydroxide	OH -1
Cyanide	CN "
Nitrate	NO <sub>3</sub> "
Nitrite	NO <sub>2</sub> 1
Chromate	CrO <sub>e</sub> <sup>-2</sup>
Oxalate	C2O4-3
Sulfate	SO2
Sulfite	SO <sub>3</sub> -2
Hydrogen sulfate	HSO <sub>4</sub> <sup>-1</sup>
Phosphate	PO <sub>4</sub> -3
Phosphite	PO <sub>3</sub> -3
Hydrogen phosphate	HPO <sub>4</sub> <sup>-2</sup>

Metric Conversions		
Umit	Abbr.	Meaning
Mega-	M	10 *
killo-	- k	10
hecto-	h	10 2
deca-	dk	10 1
Liter	L	1
gram	9	1
meter	m	1
deci-	d	10-1
coenti-	C	10 -2
milli-	m	100
milcro-	34	10 4
nano-	n	10 **

Cal	lorie Coversions	
Water	= I cal/ g*C	
1000 c	al = 1 Cal	

# **High School Chemistry Formulas Cheat Sheet**

**ML Morrison** 

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