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Equilibrium Expressions Worksheet

Look at the chemical equation $aA + bB \Leftrightarrow cC + dD$. In this reaction A, B, C, and D are chemical substances, and a, b, c, and d are the coefficients in the chemical equation. Since the equation is reversible, the reactants will not necessarily run to completion—so there won't be a limiting reactant. The forward equation and the backward equation will reach an equilibrium point where the forward equation is using A and B as fast as the backwards equation is making A and B.

If the temperature is held constant, the reaction will always reach equilibrium in such a way such that the equilibrium constant is the same. (At a different temperature the equilibrium constant for the reaction would be a constant but would be a different number.)

Equilibrium constant (Kee)

The relationship between the concentration of products and reactants at equilibrium can be expressed by an equilibrium constant K_{eq} (or K). Given the chemical equation in the form $aA + bB \Leftrightarrow cC + dD$, K_{eq} is found by the equation

K== [C]* [D]* [A]* [B]*

Note: [] denotes molar concentrations.

But when writing equilibrium expressions, reactants or products that are pure <u>solids</u> or pure <u>liquids</u> do not appear in the expression since no matter how much of it there is its density, and therefore its concentration, is always the same.

- 1. Write equilibrium expressions for each in the box provided:
 - a. $O_3(g) + NO(g) \Leftrightarrow O_2(g) + NO_2(g)$
 - b. $2CO(g) + O_2(g) \Leftrightarrow 2CO_2(g)$
 - e. $NH_4NO_3(s) \Leftrightarrow N_2O(g) + 2H_2O(l)$
 - d. $2H_2O(g) \Leftrightarrow 2H_3(g) + O_3(g)$

ь,		

c.			

Calculations involving Kos

The concentrations of the solutions/gases involved in the reversible reactions at equilibrium are known as the equilibrium concentrations. Since K_{eq} represents the system at equilibrium, only the equilibrium concentrations can be used in the equilibrium expression (<u>nov</u> the initial concentrations).

 PCl₂(g) ⇔ PCl₂(g) + Cl₂(g). What is the equilibrium constant if the equilibrium concentrations are as follows: PCl₂ is 0.0096 mol/L, PCl₃ is 0.0247 mol/L and Cl₂ is 0.0247 mol/L? Ans: 0.064

At 1000°C, a 1.00 L container has an equilibrium mixture consisting of 0.102 mol of ammonia, 1.03 mol of nitrogen, and 1.62 mol of hydrogen. Calculate the K_{eq} for the equilibrium system. N₂(g) + 3H₂ ⇔2NH₃(g) Ans: 0.00238

At a given temperature, the K_{eq} for the reaction 2HI(g) ⊕ H₂(g) + I₂(g) is 1.40 x 10⁻². If the concentration of both H₂ and I₂ at equilibrium are 2.00 x 10⁻⁴M, find the concentration of HI.
Ans: 0.00169M

Genetic Equilibrium Worksheet 16 1

Arturo Cuomo

Genetic Equilibrium Worksheet 16 1:

Decoding Genetic Equilibrium Worksheet 16 1: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "Genetic Equilibrium Worksheet 16 1," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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