```
SOQL
          Integer NUM = 10;
                                       Query
          Account[] accs;
          // Clean up old data
 Variable
          accs = [SELECT Id FROM Account WHERE name LIKE 'test%'];
Declaration
          Delete accs;
          accs = new Account[NUM];
          For (Integer I = 0; I < NUM; I++) {
 Control
              accs[I] = new Account(name='test ' + I,
 Structure
                  outstandingshares c = I);
          insert accs;
          Contact[] cons = new Contact[0];
          For (Account acc : accs) {
 Array (list)
              cons.add(new Contact(lastName = acc.name + '1',
                  accountid = acc.id));
              cons.add(new Contact(lastName = acc.name + '2',
 Data (DML)
                  accountid = acc.id));
 Operation
          Insert cons;
```

# **Forcecom Apex Code Developers Guide**

**Kevin J. Poorman** 

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web a at atmospheric pressure helium can exist in all three phases as well as a supercritical fluid phase near absolute zero choice b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure choice c

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web stant is the foundation of thermochemistry it allows us to describe heat flow in terms of the change in a state function enthalpy change is path independent note with regard to choice b that enthalpy change equals internal energy change only if the volume of the system remains constant and thus no thermody namic work is performed 5 b

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web conceptual answers in order for the reaction to occur spontaneously  $\Delta g$  for the reaction must be less than zero in this

case  $\Delta s$  must be positive and the  $t\Delta s$  term outweighs the positive value of  $\Delta h$ 

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web 1 answer a explanation the value of  $\Delta h$  is negative therefore the reaction is exothermic or heat releasing 2 answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta h$  must also be halved 3 answer c explanation option c is incorrect because 566 kj mol is the  $\Delta h$  per mole of o2 and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes h 2 o s h 2 o l h 6 01 kj mol this is an endothermic process it requires 6 01 kj to melt one mole of ice h 2 o s the enthalpy value will change if the number of moles varies from the 1 1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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