

# How to Calculate an Empirical Formula

A 283.9 g sample of unknown composition was carefully analyzed and found to consist of 49.8 g of sodium, 113.6 g of chromium, and 121.4 g of oxygen. Determine the empirical formula of the unknown compound.

① Convert mass to moles

② Divide by lowest amount of moles

$$\frac{49.8 \text{ g Na}}{22.99 \text{ g Na}} \times \frac{1 \text{ mol Na}}{1} = \frac{2.166 \text{ mol Na}}{2.166} = \boxed{1} \times 2$$

SUBSCRIPTS

$$\frac{113.6 \text{ g Cr}}{52.00 \text{ g Cr}} \times \frac{1 \text{ mol Cr}}{1} = \frac{2.185 \text{ mol Cr}}{2.166} = \boxed{1} \times 2$$

$$\frac{121.4 \text{ g O}}{16.00 \text{ g O}} \times \frac{1 \text{ mol O}}{1} = \frac{7.588 \text{ mol O}}{2.166} = \boxed{3.5} \times 2 =$$

# Empirical Formula Determination

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A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Empirical Formula Determination:**

## Reviewing **Empirical Formula Determination**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Empirical Formula Determination**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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