

CHAPTER 1

MATTER AND MEASUREMENT

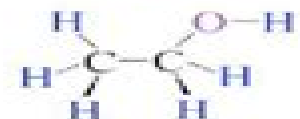
Chemistry: deals with composition, structure, and reactions of matter

- **Matter** – has mass and occupies space
 - Mass – measures quantity of matter
 - Weight – result of gravitational attraction between matter
 - Weight and mass are NOT the same
- **Composition** – what matter is made of and how much of each component is present
 - Expressed by...
 - Weight/mass
 - Volume
 - Percent
 - # of moles
 - # of atoms
 - *Macroscopic* – can be seen and weighed
 - *Submicroscopic level* – number and types of atoms
 - *Atoms* = simple units of matter
 - *Molecules* = combinations of atoms
 - Don't separate when dissolved like compounds → ionic compounds don't exist as molecules and separate when dissolved
 - *Qualitative* = ethanol consists of carbon, hydrogen and oxygen
 - *Quantitative* = ethanol = C_2H_6O
- **Structure** – arrangement of components and how they are bonded
 - Different arrangements of same combination of components = isomers
 - Structures rotate freely
 - Dashed lines = behind dark triangles = forward
- **Reactions** – changes in composition and structure
 - What products?
 - How much of product? (stoich)
 - How fast change occurs?
 - What energy changes accompany rxn?



- exothermic reaction → produces energy over a long period of time

Ethanol
 C_2H_6O
or
 C_2H_5OH



Scientific Method

- Experiment (record observations) → numerical + observational
 - Careful recordings and analysis of data
 - Controlled conditions so we pinpoint what is causing another thing to happen
 - Reproducible
- Draw a Conclusion – **LAW**
 - Concise statement about a basic relationship or regularity of nature drawn from observations

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**Edward J. Neth, Paul Flowers, Klaus
Hellmut Theopold, Richard
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Table of Contents Gradpoint Chemistry Chapter 1

1. Understanding the eBook Gradpoint Chemistry Chapter 1
 - The Rise of Digital Reading Gradpoint Chemistry Chapter 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Gradpoint Chemistry Chapter 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gradpoint Chemistry Chapter 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gradpoint Chemistry Chapter 1
 - Personalized Recommendations
 - Gradpoint Chemistry Chapter 1 User Reviews and Ratings
 - Gradpoint Chemistry Chapter 1 and Bestseller Lists
5. Accessing Gradpoint Chemistry Chapter 1 Free and Paid eBooks
 - Gradpoint Chemistry Chapter 1 Public Domain eBooks
 - Gradpoint Chemistry Chapter 1 eBook Subscription Services
 - Gradpoint Chemistry Chapter 1 Budget-Friendly Options

6. Navigating Gradpoint Chemistry Chapter 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Gradpoint Chemistry Chapter 1 Compatibility with Devices
 - Gradpoint Chemistry Chapter 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gradpoint Chemistry Chapter 1
 - Highlighting and Note-Taking Gradpoint Chemistry Chapter 1
 - Interactive Elements Gradpoint Chemistry Chapter 1
8. Staying Engaged with Gradpoint Chemistry Chapter 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gradpoint Chemistry Chapter 1
9. Balancing eBooks and Physical Books Gradpoint Chemistry Chapter 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gradpoint Chemistry Chapter 1
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gradpoint Chemistry Chapter 1
 - Setting Reading Goals Gradpoint Chemistry Chapter 1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gradpoint Chemistry Chapter 1
 - Fact-Checking eBook Content of Gradpoint Chemistry Chapter 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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