

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

- 1) How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres and a temperature of 340 K? $PV = nRT$
 $V = 120\text{ L}$
 $P = 2.3\text{ atm}$
 $T = 340\text{ K}$
 $n = ?$

$$(2.3\text{ atm})(120\text{ L}) = n(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(340\text{ K})$$

$$\boxed{n = 9.9\text{ mol}}$$
- 2) If I have a 50 liter container that holds 45 moles of gas at a temperature of 200° C, what is the pressure inside the container?
 $V = 50\text{ L}$
 $n = 45\text{ mol}$
 $T = 200^\circ\text{C} = 473\text{ K}$
 $P = ?$

$$(P)(50\text{ L}) = (45\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(473\text{ K})$$

$$\boxed{P = 30\text{ atm}} \quad \text{or} \quad \boxed{4000\text{ kPa}}$$
- 3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?
 $V = 1.0\text{ L}$
 $n = 2\text{ mol}$
 $T = 1400^\circ\text{C} = 1673\text{ K}$
 $P = ?$

$$(P)(1.0\text{ L}) = (2\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(1673\text{ K})$$

$$\boxed{P = 275\text{ atm} = 280\text{ atm}}$$

$$\text{or } 28,000\text{ kPa}$$
- 4) How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 200 atmospheres?
 $V = 30\text{ L}$
 $T = 300\text{ K}$
 $P = 200\text{ atm}$
 $n = ?$

$$(200\text{ atm})(30\text{ L}) = n(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(300\text{ K})$$

$$\boxed{n = 243.6\text{ mol} = 240\text{ mol}}$$
- 5) I have a balloon that can hold 100 liters of air. If I blow up this balloon with 3.0 moles of oxygen gas at a pressure of 1.0 atmosphere, what is the temperature of the balloon?
 $V = 100\text{ L}$
 $n = 3\text{ mol}$
 $P = 1\text{ atm}$
 $T = ?$

$$(1\text{ atm})(100\text{ L}) = (3\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(T)$$

$$\boxed{T = 406\text{ K} = 133^\circ\text{C} = 130^\circ\text{C}}$$

Ideal Gas Law Practice Work Sheet

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author names.

Ideal Gas Law Practice Work Sheet:

Holt Chemistry Ralph Thomas Myers, 2004 *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **A Guide to Teaching in the Active Learning Classroom** Paul Baepler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03

While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room's central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed. The importance of learning how to use these classrooms well and to capitalize on their special features is paramount. The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls. This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively. Among the questions this book addresses are: How can instructors mitigate the apparent lack of a central focal point in the space? What types of learning activities work well in the ALCs and take advantage of the affordances of the room? How can teachers address familiar classroom management challenges in these unfamiliar spaces? If assessment and rapid feedback are critical in active learning, how do they work in a room filled with circular tables and no central focus point? How do instructors balance group learning with the needs of the larger class? How can students be held accountable when many will necessarily have their backs facing the instructor? How can instructors evaluate the effectiveness of their teaching in these spaces? This book is intended for faculty preparing to teach in or already working in this new classroom environment for administrators planning to create ALCs or experimenting with provisionally designed rooms and for faculty developers helping teachers transition to using these new spaces.

American Journal of Physics, 2005 *Spreadsheets for Chemists* Gordon Filby, 1995

A practical guide. *Spreadsheets for Chemists* shows chemists of all levels how to use spreadsheet programs in their daily work. It highlights the possibilities provided by Lotus 1 2 3, the most widely used spreadsheet program in the sciences. Apart from hundreds of example fragments, it features a detailed discussion of the most relevant functions and all the macro commands. An accompanying diskette containing 57 worksheets involving many different fields of chemical research and teaching. An extensive glossary of spreadsheet terms. Three appendices covering 1 2 3's competitors and add-in packages, the use of Windows-based spreadsheets, and how what-if analysis and back-solving is applied. Although the disk examples were developed for Lotus 1 2 3 DOS Versions 2 x x 2 4, the worksheets are compatible with the newer Windows versions and those of their main competitors, Borland's Quattro Pro and Microsoft's Excel. Several compatible spreadsheets, AsEasyAs Proqube, might also be used as inexpensive alternatives. The author is a senior scientist at the Nuclear Research Centre in Karlsruhe, Germany. He has been using spreadsheet software for nearly ten years successfully in a variety of chemical problems. Simulations and Student

Learning Matthew Schnurr, Anna MacLeod, 2021-01-04 The book underlines the value of simulation based education as an approach that fosters authentic engagement and deep learning Safety Professional's Reference and Study Guide W. David Yates, 2017-12-12 While there are numerous technical resources available often you have to search through a plethora of them to find the information you use on a daily basis And maintaining a library suitable for a comprehensive practice can become quite costly The new edition of a bestseller Safety Professional's Reference and Study Guide Second Edition provides a single source reference that contains all the information required to handle the day to day tasks of a practicing industrial hygienist New Chapters in the Second Edition cover Behavior based safety programs Safety auditing procedures and techniques Environmental management Measuring health and safety performance OSHA's laboratory safety standard Process safety management standard BCSPs Code of Ethics The book provides a quick desk reference as well as a resource for preparations for the Associate Safety Professional ASP Certified Safety Professional CSP Occupational Health and Safety Technologist OHST and the Construction Health and Safety Technologist CHST examinations A collection of information drawn from textbooks journals and the author's more than 25 years of experience the reference provides as the title implies not just a study guide but a reference that has staying power on your library shelf **Resources in Education** ,1974
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