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# How To Do A System Restore On Toshiba Laptop

**Kathleen Armour**



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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} \rightarrow \text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} \rightarrow 2\text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add h ions or the h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride cl ions in acidic solution created by jay

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web balance each half reaction in basic solution  $4\text{Fe}^{2+} + 5\text{H}_2\text{O} + 6\text{OH}^- \rightarrow 4\text{Fe}^{3+} + 2\text{NO} + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox

reaction in acid solution using the half reaction method 8 h<sub>2</sub>o<sub>2</sub> cr<sub>2</sub>o<sub>7</sub><sup>2-</sup> 2 9 teo<sub>3</sub> 2 n<sub>2</sub>o<sub>4</sub> o<sub>2</sub> cr<sub>3</sub> te 10 reo<sub>4</sub> 11 pbo<sub>2</sub> io i<sub>2</sub> pb<sub>2</sub> 12 as no<sub>3</sub> io<sub>3</sub> re io<sub>3</sub> h<sub>2</sub>aso<sub>4</sub>

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding h<sup>+</sup> 2 o<sup>2-</sup> h<sup>+</sup> and e<sup>-</sup> in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h<sup>+</sup> 2no<sup>2-</sup> 2 c 6 h 6 o 6 2h<sup>+</sup> 2 o<sup>2-</sup> 2no choose 1 answer c 6 h 8 o 6 and no<sup>2-</sup> a c 6 h 8 o 6 and no<sup>2-</sup> no 2 and c 6 h 8 o 6 b no<sup>2-</sup> and c 6 h 8 o 6 h<sup>+</sup> and no<sup>2-</sup> c h<sup>+</sup> and

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