

8 • Electron Configurations & Periodicity

WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ s: <u>18</u>	Element: Mg # of e ⁻ s: <u>12</u>	Element: N # of e ⁻ s: <u>7</u>	Element: Li # of e ⁻ s: <u>3</u>	Element: P # of e ⁻ s: <u>15</u>	Element: Cl # of e ⁻ s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^6$
Ar $[\text{Ne}] 3s^2 3p^6$
- Mg $1s^2 2s^2 2p^6 3s^2$
Mg $[\text{Ne}] 3s^2$
- N $1s^2 2s^2 2p^3$
N $[\text{He}] 2s^2 2p^3$
- Li $1s^2 2s^1$
Li $[\text{He}] 2s^1$
- P $1s^2 2s^2 2p^6 3s^2 3p^3$
P $[\text{Ne}] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
Cl $[\text{Ne}] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



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AW Rasmussen



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Table of Contents Holt Chemistry Section Review Answers Electron Configuration

1. Understanding the eBook Holt Chemistry Section Review Answers Electron Configuration
 - The Rise of Digital Reading Holt Chemistry Section Review Answers Electron Configuration
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Section Review Answers Electron Configuration
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Section Review Answers Electron Configuration
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Section Review Answers Electron Configuration
 - Personalized Recommendations
 - Holt Chemistry Section Review Answers Electron Configuration User Reviews and Ratings
 - Holt Chemistry Section Review Answers Electron Configuration and Bestseller Lists

5. Accessing Holt Chemistry Section Review Answers Electron Configuration Free and Paid eBooks
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 - Holt Chemistry Section Review Answers Electron Configuration eBook Subscription Services
 - Holt Chemistry Section Review Answers Electron Configuration Budget-Friendly Options
6. Navigating Holt Chemistry Section Review Answers Electron Configuration eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Chemistry Section Review Answers Electron Configuration Compatibility with Devices
 - Holt Chemistry Section Review Answers Electron Configuration Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Section Review Answers Electron Configuration
 - Highlighting and Note-Taking Holt Chemistry Section Review Answers Electron Configuration
 - Interactive Elements Holt Chemistry Section Review Answers Electron Configuration
8. Staying Engaged with Holt Chemistry Section Review Answers Electron Configuration
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Section Review Answers Electron Configuration
9. Balancing eBooks and Physical Books Holt Chemistry Section Review Answers Electron Configuration
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Section Review Answers Electron Configuration
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt Chemistry Section Review Answers Electron Configuration
 - Setting Reading Goals Holt Chemistry Section Review Answers Electron Configuration
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Holt Chemistry Section Review Answers Electron Configuration
 - Fact-Checking eBook Content of Holt Chemistry Section Review Answers Electron Configuration
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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