1. Regarding the following reaction:

$$F_{2 (g)} + 2 \Gamma_{(aq)} \rightarrow 2 F_{(aq)} + I_{2 (q)}$$

- a. List the species being oxidized:_____ List the species being reduced:_____
- b. Calculate E° for this cell.
- c. Which species receives electrons from the anode?_____
- d. Which species donates electrons to the cathode?_____
- 2. Given the following half-reactions:

$$Co^{+2} + 2e^{-} \rightarrow Co$$
 $E^{0} = -0.277 \text{ V}$
 $Ce^{+4} + e^{-} \rightarrow Ce^{+3}$ $E^{0} = 1.61 \text{ V}$

- a. Write the overall equation for the galvanic cell. Be sure to balance the number of electrons.
- b. Calculate the E° for the cell.
- e. Designate which 1/2 reaction occurs...

at the anode: _____ and at the cathode: _____

- d. Describe the direction of electron flow.
- 3. Find the following using a voltaic cell with the line notation: Al₁₀[Al³⁴][Pb⁷⁴]Pb₁₀.
 - a. Calculate the E° for the reaction.
 - b. Write the balanced overall REDOX equation for the galvanic cell.
- 4. Using standard reduction potentials, calculate the cell potential (E°) for each of the following reactions:

a.
$$H_{2 (g)} + I_{2 (h)} \rightarrow 2H^*_{(eq)} + 2\Gamma_{(eq)}$$

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