

MEASUREMENT CONVERSIONS

Large → Small  Multiply them all

Small → Large  Divide and take change

Kilo	Hecto	Deca	By	Deci	Centi	Mil
kilo-	hecto-	deca-	base (L, m, g)	dec-	cent-	mill-

LENGTH

- How long an object is
- Distance between two points

feet (ft) — inch (in)

yard (yd) — mile (mi)

1 foot = 12 inches

1 yard = 3 feet

1 yard = 36 inches

1 mile = 5,280 feet

centimeter (cm) — millimeter (mm)

meter (m) — centimeter (cm)

1 centimeter = 10 millimeters

1 meter = 100 centimeters

1 meter = 1,000 millimeters

WEIGHT/MASS

- How much something weighs

How heavy an object is
pound (lb) — ounces (oz) — ton (T)

1 pound = 16 ounces

1 ton = 2,000 pounds

CAPACITY/ VOLUME

- How much a container can hold

- Finding a container

cup (c) — ounces (fl oz) — pint (pt)

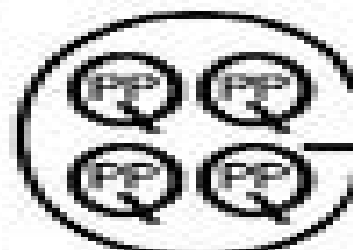
quart (qt) — gallon (gal)

1 cup = 8 ounces

1 pint = 2 cups

1 quart = 2 pints

1 gallon = 4 quarts



liter (L) — milliliter (mL)

1 liter = 1,000 milliliters

STUDY EACH NIGHT FOR 10 MINUTES & GET THIS SIGNED & DATED!

Metric Conversions Study Guide

Leo J. Malone, Theodore O. Dolter



Metric Conversions Study Guide:

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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