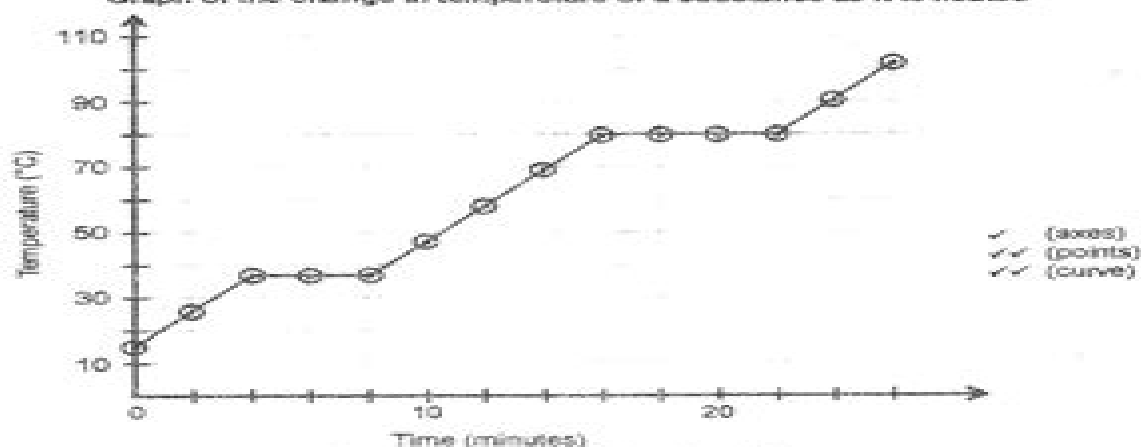


3.4

Graph of the change in temperature of a substance as it is heated

(5)



3.5 Substance exists as both a solid and a liquid ✓✓ or... is melting (2)

3.6 37 °C ✓ (2)

3.7 80 °C ✓ (1)

3.8 An indication of the average kinetic energy ✓ of the particles in a substance. ✓ (2)

3.9 Physical ✓ (1)

3.10 In the solid state the intermolecular forces are the strongest and the particles cannot move. The heater provides thermal energy to the system, which is translated into kinetic energy (increase in temperature) until the intermolecular forces between the particles can be broken. ✓✓ At this point the thermal energy is transformed into latent energy to break the bonds, so no increase in temperature occurs. ✓✓ (4)

[23]

## QUESTION 4

4.1

4.1.1 Homogeneous ✓ (1)

4.1.2 Homogeneous ✓ (1)

4.1.3 Pure substance ✓ (1)

4.2

4.2.1 Potassium bromide ✓ (2)

4.2.2 Calcium carbonate ✓ (2)

4.2.3 Ammonium chloride ✓ (2)

4.3

4.3.1  $\text{ZnSO}_4$  ✓✓ (2)4.3.2  $\text{KMnO}_4$  ✓✓ (2)4.3.3  $\text{H}_2\text{O}$  ✓✓ (2)

[17]

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## **Table of Contents March 2014 Life Sciences Question Paper And Memo**

1. Understanding the eBook March 2014 Life Sciences Question Paper And Memo
  - The Rise of Digital Reading March 2014 Life Sciences Question Paper And Memo
  - Advantages of eBooks Over Traditional Books
2. Identifying March 2014 Life Sciences Question Paper And Memo
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an March 2014 Life Sciences Question Paper And Memo
  - User-Friendly Interface
4. Exploring eBook Recommendations from March 2014 Life Sciences Question Paper And Memo

- Personalized Recommendations
  - March 2014 Life Sciences Question Paper And Memo User Reviews and Ratings
  - March 2014 Life Sciences Question Paper And Memo and Bestseller Lists
5. Accessing March 2014 Life Sciences Question Paper And Memo Free and Paid eBooks
    - March 2014 Life Sciences Question Paper And Memo Public Domain eBooks
    - March 2014 Life Sciences Question Paper And Memo eBook Subscription Services
    - March 2014 Life Sciences Question Paper And Memo Budget-Friendly Options
  6. Navigating March 2014 Life Sciences Question Paper And Memo eBook Formats
    - ePub, PDF, MOBI, and More
    - March 2014 Life Sciences Question Paper And Memo Compatibility with Devices
    - March 2014 Life Sciences Question Paper And Memo Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of March 2014 Life Sciences Question Paper And Memo
    - Highlighting and Note-Taking March 2014 Life Sciences Question Paper And Memo
    - Interactive Elements March 2014 Life Sciences Question Paper And Memo
  8. Staying Engaged with March 2014 Life Sciences Question Paper And Memo
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers March 2014 Life Sciences Question Paper And Memo
  9. Balancing eBooks and Physical Books March 2014 Life Sciences Question Paper And Memo
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection March 2014 Life Sciences Question Paper And Memo
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine March 2014 Life Sciences Question Paper And Memo
    - Setting Reading Goals March 2014 Life Sciences Question Paper And Memo
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of March 2014 Life Sciences Question Paper And Memo

- Fact-Checking eBook Content of March 2014 Life Sciences Question Paper And Memo
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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