

Mixtures and Solutions

Section 14.1 Types of Mixtures

pages 476–479

Section Assessment 14.1

page 479

- 1. Explain** Use the properties of seawater to describe the characteristics of mixtures.
Answers will vary but might include that seawater is a heterogeneous mixture with dirt and mud particles, and it is a homogeneous mixture with dissolved substances.

- 2. Distinguish** between suspensions and colloids.

Suspension particles are larger than colloidal particles. Suspension particles settle out of the mixture, whereas colloidal particles do not.

- 3. Identify** the various types of solutions. Describe the characteristics of each type of solution.

All solutions are homogeneous mixtures containing two or more substances. Solutions may be liquid, solid, or gas. Solution types are identified in Table 14.2.

- 4. Explain** Use the Tyndall effect to explain why it is more difficult to drive through fog using high beams than using low beams.

High beams are aimed farther down the road than low beams. Because the fog scatters light, there is less light from the high beams to illuminate the road than from the low beams. Also, because the high beams are aimed more directly into the fog, more of their light is reflected back toward the driver, making it more difficult to see.

- 5. Describe** different types of colloids.
See Table 14.1 for descriptions of colloid types.

- 6. Explain** Why do dispersed colloid particles stay dispersed?

The particles do not settle out because they have polar or charged layers surrounding them. These layers repel each other and prevent the particles from settling or separating.

- 7. Summarize** What causes Brownian motion?

Collisions of particles of the dispersion medium with the dispersed particles results in Brownian motion.

- 8. Compare and Contrast** Make a table that compares the properties of solutions, suspensions, and colloids.

Student tables will vary, but should include particle size, if the particles settle out, and if the particles display the Tyndall effect. A sample table follows.

Suspensions, Colloids, and Solutions			
	Particle size	Particles settle?	Tyndall effect?
Suspensions	Large (wide variation)	Yes	Yes
Colloids	1 nm–1000 nm	No	Yes
Solutions	Atomic scale (atoms, ions, and molecules)	No	No

Section 14.2 Solution Concentration

pages 480–488

Practice Problems

pages 481–488

- 9.** What is the percent by mass of NaHCO_3 in a solution containing 20.0 g NaHCO_3 dissolved in 600.0 mL H_2O ?

$$600.0 \text{ mL } \text{H}_2\text{O} \times 1.0 \text{ g/mL} = 600.0 \text{ g } \text{H}_2\text{O}$$

$$\frac{20 \text{ g } \text{NaHCO}_3}{600 \text{ g } \text{H}_2\text{O} + 20 \text{ g } \text{NaHCO}_3} \times 100 = 3\%$$

- 10.** You have 1500.0 g of a bleach solution. The percent by mass of the solute sodium hypochlorite, NaOCl , is 3.62%. How many grams of NaOCl are in the solution?

$$3.62\% = 100 \times \frac{\text{mass } \text{NaOCl}}{1500.0 \text{ g}}$$

$$\text{mass } \text{NaOCl} = 54.3 \text{ g}$$

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