

**Ions and Isotopes Worksheet**

Symbol	Name of the element	Atomic number	Protons	Neutrons	Electrons	Atomic Mass	Charge
<b>O -17</b>	<b>Oxygen</b>	<b>8</b>	<b>8</b>	<b>9</b>	<b>8</b>	<b>17</b>	<b>0</b>
<b>Xe - 131</b>	<b>Xenon</b>	<b>54</b>	<b>54</b>	<b>77</b>	<b>54</b>	<b>131</b>	<b>0</b>
$^{207}_{82}\text{Pb}$	<b>Lead</b>	<b>82</b>	<b>82</b>	<b>125</b>	<b>82</b>	<b>207</b>	<b>0</b>
<b>Na<sup>1+</sup></b>	<b>Sodium ion</b>	<b>11</b>	<b>11</b>	<b>12</b>	<b>10</b>	<b>23</b>	<b>+1</b>
<b>Sr</b>	<b>Strontium</b>	<b>38</b>	<b>38</b>	<b>50</b>	<b>38</b>	<b>88</b>	<b>0</b>
<b>Al<sup>3+</sup></b>	<b>Aluminum ion</b>	<b>13</b>	<b>13</b>	<b>14</b>	<b>10</b>	<b>27</b>	<b>+3</b>
$^{238}_{92}\text{U}$	<b>Uranium</b>	<b>92</b>	<b>92</b>	<b>146</b>	<b>92</b>	<b>238</b>	<b>0</b>
$^{75}_{33}\text{As}^{-3}$	<b>Arsenic</b>	<b>33</b>	<b>33</b>	<b>39</b>	<b>36</b>	<b>75</b>	<b>-3</b>
$^{32}_{16}\text{S}^{-2}$	<b>Sulfide ion</b>	<b>16</b>	<b>16</b>	<b>16</b>	<b>18</b>	<b>32</b>	<b>-2</b>
<b>Na</b>	<b>Sodium</b>	<b>11</b>	<b>11</b>	<b>12</b>	<b>11</b>	<b>23</b>	<b>0</b>
<b>N<sup>3-</sup></b>	<b>Nitride ion</b>	<b>7</b>	<b>7</b>	<b>7</b>	<b>10</b>	<b>14</b>	<b>-3</b>
$^{65}_{29}\text{Cu}$	<b>Copper</b>	<b>29</b>	<b>29</b>	<b>36</b>	<b>29</b>	<b>65</b>	<b>0</b>
$^{120}_{80}\text{Hg}$	<b>Mercury isotope</b>	<b>80</b>	<b>80</b>	<b>40</b>	<b>80</b>	<b>120</b>	<b>0</b>
<b>Ag</b>	<b>silver</b>	<b>47</b>	<b>47</b>	<b>61</b>	<b>47</b>	<b>108</b>	<b>0</b>
<b>Zr</b>	<b>zirconium</b>	<b>40</b>	<b>40</b>	<b>51</b>	<b>91</b>	<b>40</b>	<b>0</b>
<b>Cr</b>	<b>chromium</b>	<b>24</b>	<b>24</b>	<b>28</b>	<b>24</b>	<b>56</b>	<b>0</b>
$^{261}_{104}\text{Rf}$	<b>rutherfordium</b>	<b>104</b>	<b>104</b>	<b>158</b>	<b>104</b>	<b>261</b>	<b>0</b>

# Isotopes And Ions Worksheet Answer Key

**E Durkheim**



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